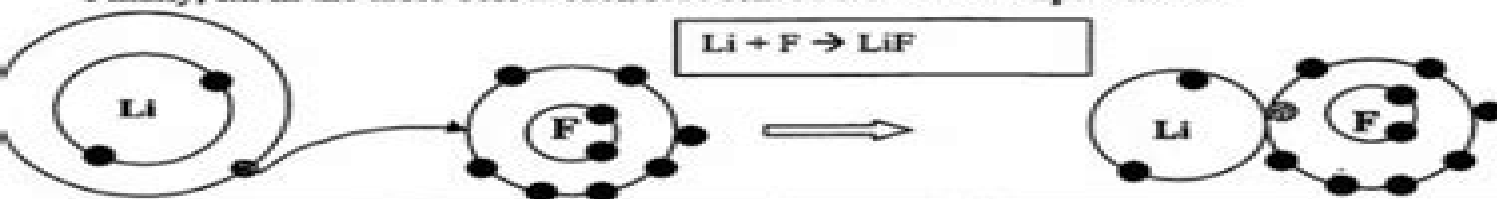



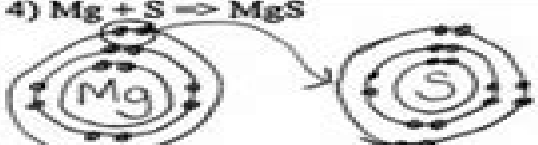



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For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1	1 to F	$\text{Li}^+$
F	7	1 from Li	$\text{F}^-$

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$	Li	1	1 to Cl	$\text{Li}^+$
	Cl	7	1 from Li	$\text{Cl}^-$
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$	Ca	2	2 to O	$\text{Ca}^{2+}$
	O	6	2 from Ca	$\text{O}^{2-}$
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$	Be	2	1 to each F	$\text{Be}^{2+}$
	F	7	Each 1 gets 1 from Be	$\text{F}^-$
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$	Mg	2	2 to S	$\text{Mg}^{2+}$
	S	6	2 from Mg	$\text{S}^{2-}$
5) $\text{K} + \text{F} \Rightarrow \text{KF}$	K	1	1 to F	$\text{K}^+$
	F	7	1 from K	$\text{F}^-$

# Ionic Covalent Bonding Worksheets With Answers

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## **Table of Contents Ionic Covalent Bonding Worksheets With Answers**

1. Understanding the eBook Ionic Covalent Bonding Worksheets With Answers
  - The Rise of Digital Reading Ionic Covalent Bonding Worksheets With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Covalent Bonding Worksheets With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Covalent Bonding Worksheets With Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Covalent Bonding Worksheets With Answers
  - Personalized Recommendations
  - Ionic Covalent Bonding Worksheets With Answers User Reviews and Ratings
  - Ionic Covalent Bonding Worksheets With Answers and Bestseller Lists
5. Accessing Ionic Covalent Bonding Worksheets With Answers Free and Paid eBooks

- Ionic Covalent Bonding Worksheets With Answers Public Domain eBooks
  - Ionic Covalent Bonding Worksheets With Answers eBook Subscription Services
  - Ionic Covalent Bonding Worksheets With Answers Budget-Friendly Options
6. Navigating Ionic Covalent Bonding Worksheets With Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Ionic Covalent Bonding Worksheets With Answers Compatibility with Devices
    - Ionic Covalent Bonding Worksheets With Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Ionic Covalent Bonding Worksheets With Answers
    - Highlighting and Note-Taking Ionic Covalent Bonding Worksheets With Answers
    - Interactive Elements Ionic Covalent Bonding Worksheets With Answers
  8. Staying Engaged with Ionic Covalent Bonding Worksheets With Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Ionic Covalent Bonding Worksheets With Answers
  9. Balancing eBooks and Physical Books Ionic Covalent Bonding Worksheets With Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Ionic Covalent Bonding Worksheets With Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Ionic Covalent Bonding Worksheets With Answers
    - Setting Reading Goals Ionic Covalent Bonding Worksheets With Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Ionic Covalent Bonding Worksheets With Answers
    - Fact-Checking eBook Content of Ionic Covalent Bonding Worksheets With Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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interaction that is responsible for forming the following solutions a kno in water b br in benzene c h c glycerol 3 2 6 6 ch oh  
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