

1. Use the VSEPR theory to predict molecular geometry of IF_5 .

a) tetrahedral	c) trigonal bipyramidal	e) octahedral
b) see-saw	d) square pyramidal	

a) bent c) trigonal pyramidal e) trigonal planar
b) tetrahedral d) T-shaped

a) 100 b) 120 c) 110 d) 90 e) 80

a) non-saw c) trigonal planar e) T-shaped
b) trigonal pyramidal d) tetrahedral

a) linear
b) bent
c) trigonal planar
d) tetrahedral
e) T shaped

a) trigonal planar c) trigonal bipyramidal e) square planar
b) tetrahedral d) octahedral

a) trigonal planar c) trigonal bipyramidal e) square planar
b) tetrahedral d) octahedral

a) 90 c) none of these
b) 109.5 d) 180

9. The species which has tetrahedral as the valence electron geometry is

a) CO_2 c) SiF_4 e) none of these
 b) SO_2 d) XeF_4

a) CO_2 c) SiF_4 e) none of these
 b) SO_2 d) XeF_4

11. The species which is polar is _____.

a) CO_2 c) SiF_4 e) none of these

b) SO_2 d) XeF_4

☐ a. SO_2
☐ b. SO_3
☐ c. SF_6
☐ d. SOF_2

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WJ Hussar



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