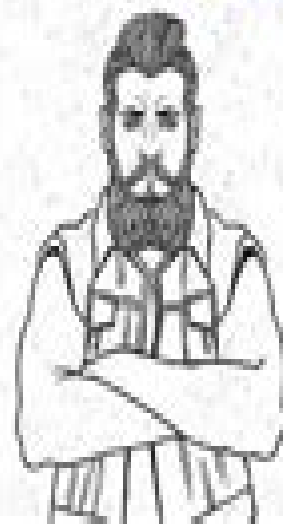


NINJA 250



CARBURETORS

STEP by STEP



Kawasaki Ninja 250r Parts Diagram

KJ Lindholm-Leary

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$$\begin{array}{l} 1 \text{ N}_2 + \text{F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} \\ 3 \text{ HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \\ \text{GaBr}_3 + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3 \text{ NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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 using the following equation $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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assuming CeCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{HBr}$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $\text{CaH}_2 + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
 $\text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{CO}_3 + \text{H}_2\text{SO}_4 + \text{H}_2\text{O}$
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same thing goes with the o3 i am so lost

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