

Objective type

- Q.1 Which of the following is not a general characteristic of equilibria involving physical processes?
- (1) Equilibrium is possible only in a closed system at a given temperature.
 - (2) All measurable properties of the system remain constant.
 - (3) All the physical processes stop at equilibrium.
 - (4) The opposing processes occur at the same rate and there is dynamic but stable condition.
- Q.2 Which of the following statements is incorrect?
- (1) In equilibrium mixture of ice and water kept in perfectly insulated flask mass of ice and water does not change with time.
 - (2) The intensity of red colour increases when oxalic acid is added to a solution containing iron (III) nitrate and potassium thiocyanate.
 - (3) On addition of catalyst the equilibrium constant value is not affected.
 - (4) Equilibrium constant for a reaction with negative ΔH value decreases as the temperature increases.
- Q.3 When hydrochloric acid is added to cobalt nitrate solution at room temperature, the following reaction takes place and the reaction mixture becomes blue. On cooling the mixture it becomes pink. On the basis of this information mark the correct answer.
- $$\underset{\text{(pink)}}{[\text{Co}(\text{H}_2\text{O})_6]^{3+}}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \underset{\text{(blue)}}{[\text{CoCl}_4]^{2-}}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$$
- (1) $\Delta H > 0$ for the reaction
 - (2) $\Delta H < 0$ for the reaction
 - (3) $\Delta H = 0$ for the reaction
 - (4) The sign of ΔH cannot be predicted on the basis of this information.
- Q.4 What will be the correct order of vapour pressure of water, acetone and ether at 30°C . Given that among these compounds, water has maximum boiling point and ether has minimum boiling point?
- (1) Water < ether < acetone
 - (2) Water < acetone < ether
 - (3) Ether < acetone < water
 - (4) Acetone < ether < water
- Q.5 In which of the following reactions, the equilibrium remains unaffected on addition of small amount of argon at constant volume?
- (1) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 - (2) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
 - (3) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
 - (4) The equilibrium will remain unaffected in all the three cases

Assertion & Reason

- Q.6 Assertion (A): In the dissociation of PCl_5 at constant pressure and temperature addition of helium at equilibrium increases the dissociation of PCl_5 .
- Reason (R) : Helium removes Cl_2 from the field of action.
- (1) Both A and R are true and R is correct explanation of A.
 - (2) Both A and R are true but R is not correct explanation of A.
 - (3) A is true but R is false.
 - (4) Both A and R are false.

Mcqs Of Environmental Chemistry

John Moore



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methods of removing hardness properties of water water pollution and waterborne diseases The chapter Hydrocarbons MCQs covers topics of alkanes alkenes and alkynes The chapter Organic Chemistry MCQs covers topics of organic compounds alcohols sources of organic compounds classification of organic compounds uses of organic compounds alkane and alkyl radicals and functional groups The chapter Atmosphere MCQs covers topics of atmosphere composition air pollutants climatology global warming meteorology ozone depletion and troposphere

ENVIRONMENTAL CHEMISTRY

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If you are muddled by the given options then give each option a true and false test. Instead of getting confused, consider all the possibilities and neglect the incorrect options. Hence, in this way, the most appropriate answer could be easily spotted.

Topic wise MCQs for NEET Chemistry: Environmental Chemistry, Organic Chemistry, P Block Elements, States of Matter, Redox Reaction, Chemical Equilibrium, Thermodynamics, Periodic Classification of Elements, s Block Elements, Chemical Bonding, Molecular Structure, Structure of Atom, Some Basic Concepts of Chemistry, Hydrocarbons, Some Basic Concepts of Chemistry, Solid State, Structure of Atom, Solutions, Classification of Elements and Periodicity in Properties, Electrochemistry, Chemical Bonding and Molecular Structure, Chemical Kinetics, States of Matter, Gases and Liquids, Surface Chemistry, Thermodynamics, General Principles and Processes of Isolation of Elements, Equilibrium, Redox Reactions, d and f Block Elements, Hydrogen, Coordination Compounds, Haloalkanes and Haloarenes, Some p Block Elements, Alcohols, Phenols and Ethers, Organic Chemistry, Some Basic Principles and Techniques, Aldehyde, Ketones and Carboxylic Acids, Hydrocarbons, Organic Compounds Containing Nitrogen, Environmental Chemistry, Biomolecules, Polymers and Chemistry in Everyday Life.

Use a step wise approach to solve conceptual and complex questions. Several times Matching type Questions are asked where the students are required to find the mismatched or the correctly matched option. Some of the questions asked in the NEET exam are entirely memory based, therefore the students are advised to memorize the common names, scientific names, concepts and important definitions. Around 40% of the questions asked in the NEET exam are application based. Therefore, students need to focus more on the concepts along with its applications in order to score well in the examination. Chemistry topics broadly can be grouped into Organic, Inorganic and Physical chemistry. Organic Chemistry is a vast section and holds the maximum weightage. Its concepts and principles are important. One needs to be conversant with properties of groups, conversions, recognizing organic compounds etc. Interlinking between different chapters from organic chemistry can be of great help. The students must primarily focus on reading NCERT textbooks. Several times the questions asked in NEET exam are taken directly from the NCERT textbooks. Initially, avoid answering those questions for which you are not confident, because your wrong answer may reduce your final score. In order to utilize your time appropriately, divide the three hours of examination time as per your comfort among Physics, Chemistry and Biology. Initially, focus on attempting all easy questions and later on pick the difficult ones. By this way, your confidence will be elevated and you will also get more time to answer hard questions. Practice previous years' question papers, mock tests and sample papers to get an idea on how to answer MCQ questions efficiently. Preparing at an early stage is what an MCQ exam requires. Avoid guesswork for negative marking questions as they might lower your final score. These tips can be very helpful for students to answer difficult and brain teaser questions. Prior preparations and practice are mandatory aspects of any examination. Hence, to crack highly competitive examination like NEET, it is mandatory for students to prepare well and acquire the skills to tackle multiple choice questions effectively. Rather than just following mere guesswork, the aspirants can focus on the tips discussed to tackle Multiple Choice Questions.

in NEET in the right manner **Environmental Chemistry** John Moore,2012-12-02 Environmental Chemistry concerns with the broad interpretation on what environmental chemistry is and discusses chemistry in relation to environmental topics The book is divided into seven parts Part I discusses the origins of different elements and interstellar molecules the development of the earth and the chemical evolution of life Part II talks about energy and its theoretical treatment the origin development and problems related to fossil fuels and the developing energy sources including storage distribution and conservation Part III discusses the air the structure and properties of the atmosphere and air pollution in relation to different industries and transportation Mineral resources and solid wastes are tackled in Part IV and the principles and treatment of water are explained in Part V Part VI discusses the sustenance of life amino acids and the control of toxins and Part VII studies the relationship of science ethics and ecology The text is good for those in the field of chemistry and wish to understand the importance of their field to the environment and for environmentalists and ecologists who want to know the relationship of chemistry with their studies **Environmental Chemistry** , *Understanding Our Environment* Roy M. Harrison,Royal Society of Chemistry (Great Britain),1999 This introductory text is aimed at those having little background knowledge of the field Developing a more international approach it emphasises links between atmosphere water and earth

Annual Report - University of the Witwatersrand University of the Witwatersrand, *Key Concepts in Environmental Chemistry* Grady Hanrahan,2011-08-02 Key Concepts in Environmental Chemistry provides a modern and concise introduction to environmental chemistry principles and the dynamic nature of environmental systems It offers an intense one semester examination of selected concepts encountered in this field of study and provides integrated tools in explaining complex chemical problems of environmental importance Principles typically covered in more comprehensive textbooks are well integrated into general chapter topics and application areas The goal of this textbook is to provide students with a valuable resource for learning the basic concepts of environmental chemistry from an easy to follow condensed application and inquiry based perspective Additional statistical sampling modeling and data analysis concepts and exercises will be introduced for greater understanding of the underlying processes of complex environmental systems and fundamental chemical principles Each chapter will have problem oriented exercises with examples throughout the body of the chapter that stress the important concepts covered and research applications case studies from experts in the field Research applications will be directly tied to theoretical concepts covered in the chapter Overall this text provides a condensed and integrated tool for student learning and covers key concepts in the rapidly developing field of environmental chemistry Intense one semester approach to learning Application based approach to learning theoretical concepts In depth analysis of field based and in situ analytical techniques Introduction to environmental modeling **Environmental Chemistry in Society** James M. Beard,2008-09-29 Taking a nonmathematical approach to the material Environmental Chemistry in Society presents the chemistry of the environment in a way accessible to students who have little or no science background It relates

the fundamentals of chemistry to contemporary environmental issues Shows the Relevance of Chemistry in the Environment
Requiring no prio **Environmental Chemistry in the Secondary School** Gerard W. Vollmer,1974 *Environmental
Chemistry* John Wright,2004-03 Offers an accessible introduction to chemical principles and concepts and makes the subject
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The book delves into Mcqs Of Environmental Chemistry. Mcqs Of Environmental Chemistry is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Mcqs Of Environmental Chemistry, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 2: Essential Elements of Mcqs Of Environmental Chemistry
 - Chapter 3: Mcqs Of Environmental Chemistry in Everyday Life
 - Chapter 4: Mcqs Of Environmental Chemistry in Specific Contexts
 - Chapter 5: Conclusion
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