

- Directions:**
 1. An element may be assigned a positive or negative oxidation number of zero.
 2. An atom with a higher electronegativity is assigned a negative oxidation number.
 3. The sum of all oxidation numbers for a molecule must be zero.

Ion Practice Set

1. Write an equation that shows how the element is oxidized or reduced.
 2. Write the oxidation number next to the element.
 3. Write the oxidation number next to the element.

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	CHARGE OF ION	TYPE OF ION
1	Sodium	Na^+	11	10	+1	cation
2	Chlorine	Cl^-	17	18	-1	anion
3	Strontium	Sr^{2+}	38	36	+2	cation
4	Phosphorus	P^{3-}	15	18	-3	anion
5	Calcium	Ca^{2+}	20	18	+2	cation
6	Barium	Ba^{2+}	56	54	+2	cation
7	Hydrogen	H^+	1	0	+1	cation
8	Lithium	Li^+	3	2	+1	cation
9	Magnesium	Mg^{2+}	12	10	+2	cation
10	Sulfur	S^{2-}	16	18	-2	anion
11	Aluminum	Al^{3+}	13	10	+3	cation
12	Oxygen	O^{2-}	8	10	-2	anion
13	Boron	B^{3+}	5	2	+3	cation
14	Fluorine	F^-	9	10	-1	anion
15	Carbon	C^{4+}	6	2	+4	cation
16	Chlorine	Cl^-	17	18	-1	anion

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SA Adler



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