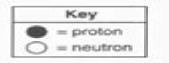
The diagram below represents the nucleus of an atom.





What are the atomic number and mass number of this atom?

- A) The atomic number is 9 and the mass number is 19.
- B) The atomic number is 9 and the mass number is 20.
- C) The atomic number is 11 and the mass number is 19.
- D) The atomic number is 11 and the mass number is 20.
- 2. Which statement concerning elements is true?
  - A) Different elements must have different numbers of isotopes.
  - Different elements must have different numbers of neutrons.
  - C) All atoms of a given element must have the same mass number.
  - D) All atoms of a given element must have the same atomic number.
- 3. Which subatomic particles are located in the nucleus of an He-4 atom?
  - A) electrons and neutrons
  - B) electrons and protons
  - C) neutrons and protons
  - D) neutrons, protons, and electrons
- 4. Which part of a helium atom is positively charged?
  - A) electron

B) neutron

C) nucleus

- D) orbital
- 5. Which total mass is the smallest?
  - A) the mass of 2 electrons
  - B) the mass of 2 neutrons
  - C) the mass of 1 electron plus the mass of 1 proton
  - D) the mass of 1 neutron plus the mass of 1 electron

- Subatomic particles can usually pass undeflected through an atom because the volume of an atom is composed of
  - A) an uncharged nucleus
  - B) largely empty space
  - C) neutrons
  - D) protons
- An experiment in which alpha particles were used to bombard thin sheets of gold foil led to the conclusion that an atom is composed mostly of
  - A) empty space and has a small, negatively charged nucleus
  - B) empty space and has a small, positively charged nucleus
  - C) a large, dense, positively charged nucleus
  - D) a large, dense, negatively charged nucleus
- 8. Which conclusion is based on the "gold foil experiment" and the resulting model of the atom?
  - A) An atom is mainly empty space, and the nucleus has a positive charge.
  - B) An atom is mainly empty space, and the nucleus has a negative charge.
  - C) An atom has hardly any empty space, and the nucleus has a positive charge.
  - D) An atom has hardly any empty space, and the nucleus has a negative charge.
- Which statement about the mass of an electron is correct?
  - A) The mass of an electron is equal to the mass of a proton.
  - B) The mass of an electron is less than the mass of a proton.
  - C) The mass of an electron is equal to the mass of a neutron.
  - D) The mass of an electron is greater than the mass of a neutron.
- 10. Experiments with gold foil indicated that atoms
  - A) usually have a uniform distribution of positive charges
  - B) usually have a uniform distribution of negative charges
  - C) contain a positively charged, dense center
  - D) contain a negatively charged, dense center

# **Mcq Pharm Chem Periodic Table**

**Rachel Sandford** 

Mcq Pharm Chem Periodic Table:

#### Reviewing Mcq Pharm Chem Periodic Table: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Mcq Pharm Chem Periodic Table**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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