

63. The metal that cannot be obtained by electrolysis of an aqueous solution of its salts is :

- (1) Cu (2) Cr (3) Ag (4) Ca

63. (4)

Factual

64. Resistance of 0.2 M solution of an electrolyte is 50  $\Omega$ . The specific conductance of the solution is  $1.4 \text{ S m}^{-1}$ . The resistance of 0.5 M solution of the same electrolyte is 200  $\Omega$ . The molar conductivity of 0.5 M solution of the electrolyte in  $\text{S m}^2 \text{ mol}^{-1}$  is :

- (1)  $5 \times 10^7$  (2)  $5 \times 10^2$  (3)  $5 \times 10^{-4}$  (4)  $5 \times 10^{-5}$

64. (3)

$$\Lambda_m = \frac{k (\text{S m})}{1000 \times \text{Molarity} (\text{mol l}^{-1})} = \frac{1}{4 \times 1000 \times 0.5} \\ = 5 \times 10^{-4} \text{ S m}^2 \text{ mol}^{-1}$$

65. CaCl crystallises in body centred cubic lattice. If 'a' is its edge length then which of the following expressions is correct ?

- (1)  $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \frac{\sqrt{3}}{2}a$  (2)  $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \frac{\sqrt{3}}{4}a$  (3)  $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \frac{3a}{4}$  (4)  $r_{\text{Ca}^{2+}} + r_{\text{Cl}^{-}} = \frac{3a}{2}$

65. (1)

Factual

66. Consider separate solutions of 0.500 M  $\text{C}_2\text{H}_5\text{OH(aq)}$ , 0.100 M  $\text{Mg}_3(\text{PO}_4)_2(\text{aq})$ , 0.250 M  $\text{KBr(aq)}$  and 0.125 M  $\text{Na}_3\text{PO}_4(\text{aq})$  at  $25^\circ\text{C}$ . Which statement is true about these solutions, assuming all salts to be strong electrolytes ?

- (1) 0.125 M  $\text{Na}_3\text{PO}_4(\text{aq})$  has the highest osmotic pressure.  
(2) 0.500 M  $\text{C}_2\text{H}_5\text{OH(aq)}$  has the highest osmotic pressure.  
(3) They all have the same osmotic pressure.  
(4) 0.100 M  $\text{Mg}_3(\text{PO}_4)_2(\text{aq})$  has the highest osmotic pressure.

66. (3)

All solutions have same osmotic pressure.

67. In which of the following reactions  $\text{H}_2\text{O}_2$  acts as a reducing agent ?

- (a)  $\text{H}_2\text{O}_2 + 2\text{H}^{+} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O}$  (b)  $\text{H}_2\text{O}_2 + 2\text{e}^{-} \rightarrow \text{O}_2 + 2\text{H}^{+}$   
(c)  $\text{H}_2\text{O}_2 + 2\text{e}^{-} \rightarrow 2\text{OH}^{-}$  (d)  $\text{H}_2\text{O}_2 + 2\text{OH}^{-} + 2\text{e}^{-} \rightarrow \text{O}_2 + 2\text{H}_2\text{O}$   
(1) (a), (c) (2) (b), (d) (3) (a), (b) (4) (c), (d)

67. (2)



68. In  $\text{S}_{\text{N}}2$  reactions, the correct order of reactivity for the following compounds :

- $\text{CH}_3\text{Cl}$ ,  $\text{CH}_3\text{CH}_2\text{Cl}$ ,  $(\text{CH}_3)_2\text{CHCl}$  and  $(\text{CH}_3)_3\text{CCl}$  is :  
(1)  $\text{CH}_3\text{CH}_2\text{Cl} > \text{CH}_3\text{Cl} > (\text{CH}_3)_2\text{CHCl} > (\text{CH}_3)_3\text{CCl}$   
(2)  $(\text{CH}_3)_2\text{CHCl} > \text{CH}_3\text{CH}_2\text{Cl} > \text{CH}_3\text{Cl} > (\text{CH}_3)_3\text{CCl}$   
(3)  $\text{CH}_3\text{Cl} > (\text{CH}_3)_2\text{CHCl} > \text{CH}_3\text{CH}_2\text{Cl} > (\text{CH}_3)_3\text{CCl}$   
(4)  $\text{CH}_3\text{Cl} > \text{CH}_3\text{CH}_2\text{Cl} > (\text{CH}_3)_2\text{CHCl} > (\text{CH}_3)_3\text{CCl}$

68. (4)

Order of reactivity of alkyl halide towards  $\text{S}_{\text{N}}2$  reaction is  $\text{CH}_3\text{X} > 1^\circ > 2^\circ > 3^\circ$

69. The octahedral complex of a metal ion  $\text{M}^{3+}$  with four monodentate ligands  $\text{L}_4$ ,  $\text{L}_3$ ,  $\text{L}_2$  and  $\text{L}_1$  absorb wavelengths in the region of red, green, yellow and blue, respectively. The increasing order of ligand strength of the four ligands is :

- (1)  $\text{L}_1 < \text{L}_2 < \text{L}_4 < \text{L}_3$  (2)  $\text{L}_1 < \text{L}_2 < \text{L}_4 < \text{L}_3$  (3)  $\text{L}_4 < \text{L}_3 < \text{L}_2 < \text{L}_1$  (4)  $\text{L}_3 < \text{L}_1 < \text{L}_2 < \text{L}_4$

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