

Matter—Properties and Changes

Section 3.1 Properties of Matter

pages 70–75

Problem-Solving Lab

1. Explain why the flow of a compressed gas must be controlled for practical and safe use.

The flow of compressed gas must be controlled to control the amount and the rate at which gas is released.

2. Predict what would happen if the valve on a full tank of compressed gas were suddenly opened all the way or if the tank were accidentally punctured.

Without the regulator device, the gas would rush out of the tank with a force powerful enough to transform the tank into a dangerous, uncontrolled projectile.

Section 3.1 Assessment

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1. Create a table that describes the three common states of matter in terms of their shape, volume, and compressibility.

	Volume	Shape	Compressibility
Solid	Definite	Definite	Incompressible
Liquid	Definite	Takes shape of container and fills container to the extent of its own volume	Virtually incompressible
Gas	Fills volume of container	Takes shape of container	Compressible

2. Describe the characteristics that identify a sample of matter as being a substance.

The sample of matter must have a uniform and unchanging composition to be a substance.

3. Classify each of the following as a physical or chemical property.

- a. Iron and oxygen form rust.
chemical

- b. Iron is more dense than aluminum.

physical

- c. Magnesium burns brightly when ignited.

chemical

- d. Oil and water do not mix.

physical

- e. Mercury melts at -39°C .

physical

4. **Organize** Create a chart that compares physical and chemical properties. Give two examples for each type of property.

The chart should make clear that physical properties can be observed without changing the composition of the sample, which is not the case for chemical properties. Mass and density are examples of physical properties. Fermentation and rusting are examples of chemical properties.

Section 3.2 Changes in Matter

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Practice Problems

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5. Use the data in the table to answer the following questions.

Aluminum and Liquid Bromine Reaction

	Before Reaction	After Reaction
Aluminum	10.3 g	0.0 g
Liquid bromine	100.0 g	8.5 g
Compound	0.0 g	

How many grams of bromine reacted? How many grams of compound were formed?

amount of bromine that reacted =
 $100.0\text{ g} - 8.5\text{ g} = 91.5\text{ g}$

amount of compound formed =
 $100.0\text{ g} + 10.3\text{ g} - 8.5\text{ g} = 101.8\text{ g}$

Manual Chemistry Matter And Change Answer Key

Daniela Niemeyer



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Quanta, Matter and Change: A Molecular Approach to Physical Change Peter Atkins, Julio de Paula, Ron Friedman, 2008-11-15 Beginning with quantum mechanics introducing statistical mechanics and progressing through to thermodynamics this new text for the two semester physical chemistry course features a wealth of new applications and insights as well as new Mathematical Background inter chapters to help students review key quantitative concepts This is a splendid book True to the authors philosophy as outlined in the preface it approaches physical chemistry by first developing the quantum theory of molecular electronic structure then by statistical arguments moves into thermodynamics and thence to kinetics Peter Taylor Review in Chemistry World Royal Society of Chemistry July 31 2009 Chemistry Laurel Dingrando, McGraw-Hill Staff, Glencoe/McGraw-Hill, 2001-08-01 **Chemistry** McGraw-Hill Staff, 2001-07 **The British National Bibliography** Arthur James Wells, 2000 Watts' Manual of Chemistry, Theoretical and Practical (based on Fownes' Manual).: Chemistry of organic carbon-compounds; or, Organic compounds Henry Watts, 1886 *El-Hi Textbooks in Print*, 1981 **Chemistry** Nicholas Drapela, Alison Hyslop, 2005 Books and Pamphlets, Including Serials and Contributions to Periodicals Library of Congress. Copyright Office, 1968 **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science**, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Cumulative Book Index**, 1925 A world list of books in the English language The United States Catalog, 1925 *Chemistry Matter and Its Changes Student Solutions Manual* James E. Brady, **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1976 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Chemical News and Journal of Industrial Science*, 1877 **Chemical News and Journal of Physical Science**, 1888 **The Chemical News and Journal of Physical Science**, 1913 El-Hi Textbooks and Serials in Print, 2006 **The Volume Library** Henry Woldmar Ruoff, 1911 **Resources in Education**, 1989

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