

(ii) What would you observe during **reaction 1**?

.....
 [1]

(iii) In **reaction 1**, compound **E** was heated under reflux with excess $\text{K}_2\text{Cr}_2\text{O}_7/\text{H}_2\text{SO}_4$.

Suggest why these conditions were used, rather than the reaction mixture being distilled during the process.

.....

 [1]

(iv) Name the type of reaction taking place in **reaction 2**.

..... [1]

(b) Describe a chemical test that you could use to detect the presence of a carbonyl group in an organic compound.

Reagent

Observation

.....
 [2]

(c) Compound **E** is a single stereoisomer.

(i) Draw the skeletal formula of one other stereoisomer of compound **E** and state the type of stereoisomerism.

type of stereoisomerism [2]

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the
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nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a
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