Question 1

- 1(a)(iii) 247 g per 100 cm3.
- 1(a)(iv) Mass of KI dissolved in 100 cm³ at 70 °C = 247 g. Mass of Ki dissolved in 100 cm³ at 30 °C = 180 g.

Mass of KI that would be precipitated out from solution on cooling from 70 °C to 30 °C = (247 - 180) g = 67 g.

- 1(a)(v) Molar mass of KI = (127 + 39) g mol⁻¹ = 166 g mol⁻¹. Moles of KI in 180 g KI = (180/166) mol = 1.08 mol. Hence moles of KI in 100 cm³ of saturated solution at 30 °C = 1.08 mol. Moles of KI in 1000 cm³ = (1.08 × 10) = 10.8 mol. Therefore concentration of saturated solution of KI at 30 °C = 10.8 mol dm⁻³.
- 1(b) KI is an ionic solid in which cations and anions are held together by ionic bonds. Water is a polar covalent solvent so it can cause the ions in KI to dissociate. Ethanol is also a polar covalent solvent but it is not as polar as water.
- 1(c)(i) Distilled water, stirring rod, beakers, aqueous barium chloride solution, filter funnel, filter paper, Bunsen burner, tripod stand, wire gauze and evaporating dish.
- 1(c)(ii) 1. Dissolve the solid mixture in distilled water.
 - Add aqueous barium chloride solution slowly with stirring until no more precipitate forms.
 - 3. Filter of the precipitate of barium sulfate.
 - Evaporate the filtrate to obtain solid NaCl.
- 1(c)(iii) Colourless solution forms when salt mixture dissolves.
 - White precipitate forms when aqueous barium chloride is added.
 - · Colourless filtrate is collected upon filtration.
 - · White solid forms upon evaporation of filtrate.
- 1(d)(i) Inference NO₂ gas formed.
 - · Acidic gas present; NO, ions present.
- 1(d)(ii) Inference Pb²⁺ ions present.
 - Pb²⁺_(aq) + Γ_(aq) → PbI_{2(q)}.

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