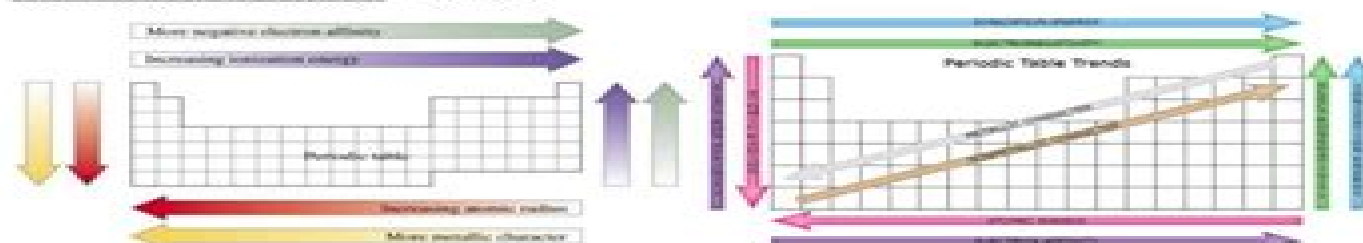


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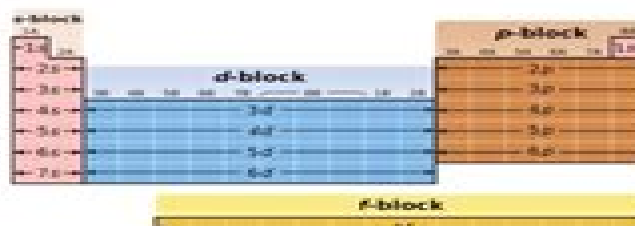
### The Nature of Matter - Trends



### Atomic Radius:

- Moving down  $\rightarrow$  size increases
  - Why? Because the # of energy shells increases and a new inner energy level is added between the nucleus and the atom's outer electron cloud
    - With each additionally inner level, the outer cloud is more shielded from the attractive force of the nucleus
    - *Electron Cloud:* a model of the atom that states we cannot know where electrons are at all times, but they are more likely to remain in certain areas (cloud) of the atom; think, vague
    - *What Affects EC Size?:* Good question! An increase in # of energy levels increases the probability that the outer electrons spend time away from the nucleus, and so the size increases
- Moving from left to right  $\rightarrow$  size decreases
  - Why? Because  $Z_{eff}$  increases and the force of attraction between the nucleus and the valence electrons is stronger (thus, the size decreases)
    - *What is  $Z_{eff}$ ?* the effective nuclear charge:  $Z$  (# of protons (atomic #)) -  $S$  (# of core electrons (shielding electrons)); conceptually, the amount of charge "felt" or "seen" by outer valence electrons
- Neutral atoms tend to get bigger as you go down a group and decrease as you go across a period
- When a neutral atom gains an electron (forming an anion), the atomic radius increases; When a neutral atom loses an electron (forming a cation), the atomic radius decreases

### Ionic Radius:



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1. Understanding the eBook National Paper Grade 11 Chemistry Final Exam 2014
  - The Rise of Digital Reading National Paper Grade 11 Chemistry Final Exam 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying National Paper Grade 11 Chemistry Final Exam 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an National Paper Grade 11 Chemistry Final Exam 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from National Paper Grade 11 Chemistry Final Exam 2014
  - Personalized Recommendations
  - National Paper Grade 11 Chemistry Final Exam 2014 User Reviews and Ratings
  - National Paper Grade 11 Chemistry Final Exam 2014 and Bestseller Lists
5. Accessing National Paper Grade 11 Chemistry Final Exam 2014 Free and Paid eBooks

- National Paper Grade 11 Chemistry Final Exam 2014 Public Domain eBooks
  - National Paper Grade 11 Chemistry Final Exam 2014 eBook Subscription Services
  - National Paper Grade 11 Chemistry Final Exam 2014 Budget-Friendly Options
6. Navigating National Paper Grade 11 Chemistry Final Exam 2014 eBook Formats
- ePub, PDF, MOBI, and More
  - National Paper Grade 11 Chemistry Final Exam 2014 Compatibility with Devices
  - National Paper Grade 11 Chemistry Final Exam 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of National Paper Grade 11 Chemistry Final Exam 2014
  - Highlighting and Note-Taking National Paper Grade 11 Chemistry Final Exam 2014
  - Interactive Elements National Paper Grade 11 Chemistry Final Exam 2014
8. Staying Engaged with National Paper Grade 11 Chemistry Final Exam 2014
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers National Paper Grade 11 Chemistry Final Exam 2014
9. Balancing eBooks and Physical Books National Paper Grade 11 Chemistry Final Exam 2014
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection National Paper Grade 11 Chemistry Final Exam 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine National Paper Grade 11 Chemistry Final Exam 2014
- Setting Reading Goals National Paper Grade 11 Chemistry Final Exam 2014
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of National Paper Grade 11 Chemistry Final Exam 2014
- Fact-Checking eBook Content of National Paper Grade 11 Chemistry Final Exam 2014
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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