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**Jianjun Gao**



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web sep 22 2021 the equilibrium system can be written in the general form  $\text{ce} + \text{ha} \rightleftharpoons \text{h} + \text{aq} + \text{aq}$  label 7 the equilibrium constant expression for this reaction is  $K = \frac{[\text{ce}][\text{h}]}{[\text{ce}][\text{ha}]}$  label 8 where we denote the equilibrium constant  $K$  with a subscript a for acid

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web equilibrium moles  $\text{hscn}$   $2.00 \times 10^{-5} \text{ mol}$   $3.00 \times 10^{-6} \text{ mol}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{hscn}$  knowing the number of moles of  $\text{Fe}^{3+}$  and  $\text{hscn}$  present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species  $\text{mol Fe}^{3+}$   $1.70 \times 10^{-5} \text{ mol}$   $\text{Fe}^{3+}$   $8.50 \times 10^{-4} \text{ M}$

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