

<u>Lewis Structure</u>	<u>Molecular Geometry</u>	<u>Bond Polarity</u>	Polar or <u>Non-polar</u>
① SiO_2 $\ddot{\text{O}}=\text{Si}=\ddot{\text{O}}$	Linear	Covalent Bond	Non-polar
② BH_3 $\begin{array}{c} \text{H} - \text{B} - \text{H} \\ \\ \text{H} \end{array}$	Trigonal planar	Covalent Bond	Non-polar
③ NH_3 $\begin{array}{c} \text{H} - \ddot{\text{N}} - \text{H} \\ \\ \text{H} \end{array}$	Trigonal Pyramidal	Covalent Bond	Polar
④ H_2O $\begin{array}{c} \cdot\cdot\text{O}\cdot\cdot \\ \diagup \quad \diagdown \\ \text{H} \quad \text{H} \end{array}$	Tetrahedral	Polar Covalent	Polar
⑤ CO $:\text{C}\equiv\text{O}:$	Linear	Covalent Bond	Polar

Lab 2models Molecular Compounds Answer

Michael Brown



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