

Lab: Single and Double Replacement Reactions

Purpose: To predict products and observe various single and double replacement reactions.

Hypotheses: Single Replacement Reactions: Use the **activity series** (p.268 - Honens) and your understanding of single replacement reactions to predict products for the following reactions. If no reaction will occur, then write "no rxn" on the product side of the equation. Remember to **BALANCE** the equation afterwards.

- 1) $\text{HCl} + \text{Mg} \rightarrow$ _____
- 2) $\text{HCl} + \text{Cu} \rightarrow$ _____ (hint: copper could be +2 in this compound)
- 3) $\text{HCl} + \text{Fe} \rightarrow$ _____ (hint: iron will be +3 in this compound)
- 4) $\text{HCl} + \text{Zn} \rightarrow$ _____

Hypotheses: Double Replacement Reactions: Use the **solubility chart** and your understanding of double replacement reactions to predict products for the following reactions (see Table 2 on the back of this paper...we're only fully writing out equations for 4 of the 16 you will test). If both of the possible products are aqueous, then write "no rxn" on the product side of the equation. Otherwise, clearly state which product will be aqueous and which will form a solid precipitate. Remember to **BALANCE** the equation afterwards.

- 1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$ _____
- 6) $\text{NaOH}(\text{aq}) + \text{Fe}(\text{NO}_3)_3(\text{aq}) \rightarrow$ _____
- 11) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$ _____
- 16) $\text{MgSO}_4(\text{aq}) + \text{Zn}(\text{NO}_3)_2(\text{aq}) \rightarrow$ _____

Procedures:

1. For each single replacement reaction, place a sample of each metal in your well plate and then place 5-6 drops of HCl (aq) on top of it. Record your observations.
2. For each double replacement reaction, place two drops of each solution carefully on the transparency, covering each "X" appropriately over your "Super High-Tech Patented Reaction Surface." Record your observations.
3. If a precipitate formed, clearly write the formula for the precipitate on the line provided.
4. Clean up your workstation according to your instructor's directions and answer all questions.

Observations/Data:

Table 1: Single Replacement Reaction Observations

	HCl (aq)
Mg	1 Observations:
Cu	2 Observations:
Fe	3 Observations:
Zn	4 Observations:

Lab 9 2 Single Replacement Reactions

Robert Bruce Thompson

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the name bar.

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Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The *Illustrated Guide to Home Chemistry Experiments* steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier's Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos *Illustrated Guide to Home Chemistry Experiments* offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry

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