

Equilibrium Lab

Instructions: Complete all sections of this report and answer all questions in complete sentences for full credit.

Title: Equilibrium Lab

Objective(s): How increasing or decreasing the concentration of one component of a reaction affects its equilibrium?

Hypothesis: *Complete the following statements for your hypotheses. Choose one of the options in brackets by highlighting, bolding, or placing your answer in a different color.*

If the amount of reactant increases, the rate of the forward reaction will [**increase** or decrease] and the rate of the reverse reaction will [increase or **decrease**].

If the amount of reactant decreases, the rate of the forward reaction will [increase or **decrease**] and the rate of the reverse reaction will [**increase** or decrease].

If the amount of product increases, the rate of the forward reaction will [increase or **decrease**] and the rate of the reverse reaction will [**increase** or decrease].

If the amount of product decreases, the rate of the forward reaction will [**increase** or decrease] and the rate of the reverse reaction will [increase or **decrease**].

Materials: For this lab, you are not expected to list the materials. But still notice the materials (lab equipment and chemicals) used in the lab to help you analyze the data, write the conclusion, and answer the post-lab questions.

Variables: *List variables below as you work through the simulation.*

Type	Variables
Controlled	
Independent	Concentration of reactants and products
Dependent	Reaction rates

Procedure: *Follow the steps described in the virtual lab. Record your observations and interpretations below.*

Equilibrium reaction in the lab:



Lab 3disturbing Equilibrium Answers

Sebastian Brünink

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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