- The number of σ bonds in N≡N is
  - a. 1 b. 2
  - e. 3
  - d. 4
- The elements in an ionic compound are held together by
  - electrostatic forces of attraction.
  - b. van der Waals forces.
  - e. the spin of paired electrons.
  - d. the formation of hybrid orbitals.
  - e. an electron pair.
- In every electrolytic and galvanic (voltaic) cell the anode is that electrode
  - at which exidation occurs.
  - b. which attracts cations.
  - at which electrons are supplied to the solution.
  - at which reduction occurs.
- 28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
  - a. 47
  - b. 62
  - c. 93d. 186
- 29. In a galvanic (voltaic) cell in which the reaction is Cd + Cu<sup>2+</sup> → Cu + Cd<sup>2+</sup> and the ions are at unit concentration (activity), the cell potential is
  - $Cd \rightarrow Cd^{2+} + 2e$  0.4021 V  $Cu \rightarrow Cu^{2+} + 2e$  0.344 V
  - a. 0.1383 V
  - b. 0.4021 V
     c. 0.344 V
  - d. 0.7461 V
  - e. 0.3677 V
- 30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
  - a.  $CaCO_2(s) \leftarrow \rightarrow CaO(s) + CO_2(g)$
  - b.  $H_2(g) + Cl_2(g) \leftrightarrow 2HCl(g)$
  - c.  $2NO(g) + O_2(g) \leftrightarrow 2NO_2(g)$
  - d. COCl₂(g) ←→ CO(g) + Cl₂(g)

Standard Potentials	E
$Mg \rightarrow Mg^2 + 2e$	2.37V
$Al \rightarrow Al^{5*}+3e$	1.66V
$Zn \rightarrow Zn^{2*}+2c$	0.76V
Fe⇒Fe <sup>2*</sup> +2e	0.44V
Cu → Cu <sup>2</sup> *+2e	0.34V
Ag→ Ag'+e	-0.80V

- Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
  - Mg and Ag
  - b. Mg and Fe
  - c. Zn and Cu
  - el. Al and Ag
  - e. Mg and Al
- 32.  $E = E^{\circ} 0.059/n \log Q$  (Nernst equation)  $[H^{\circ}] = 1.0 \text{ M initially}, P_{o2} = 1.0 \text{ atm}$

4c + O<sub>2</sub>(g)+4H\*(aq)←→ 2H<sub>2</sub>O(l) E\*=1.23V Based on the information above, which statement is correct?

- a. n = 1, since one mole of oxygen is being considered.
- Addition of base should result in an E value, which is less than 1.23 V.
- c. E is independent of the pH of the solution.
- d.  $Q = [H_2O]^2$  $[O_2][H^*]$
- 33. The equilibrium constant for the gaseous reaction C + D ← → E + 2F is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
  - a. proceed at equal rates in both directions.
  - b. proceed more rapidly to form E and F.
  - proceed more rapidly to form C and D.
  - not occur in either direction.

Compound	ΔG° <sub>ε</sub> kJ/mol	
H <sub>2</sub> O(l)	-237	
H <sub>2</sub> O(g)	-229	

- At 298 K the equilibrium constant for H<sub>2</sub>(g) + ½ O<sub>2</sub>(g) ← → H<sub>2</sub>O(I)
  - is larger than the Keq for H<sub>2</sub>(g) + ½ O<sub>2</sub>(g) ← → H<sub>2</sub>O(g)
  - b. will have a value of 1 0 at equilibrium.
  - c. cannot be computed since data on O2 and

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**Wolfgang Guggemos** 

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