

## ELECTRON CONFIGURATIONS OF ATOMS AND IONS

### OBJECTIVES

- Observe colors of various chemical solutions
- Write electron configurations for various metals, nonmetals, cations, and anions
- Relate the presence of color in an ionic solution as a characteristic of electron configurations

### PURPOSE

In this experiment, you will observe a variety of chemical solutions containing common cations and anions. You will write electron configurations for many of the ions contained in the solutions. You will observe the colored solutions and draw conclusions about the electron configurations of the metal ions in the colored solutions.

### EXPERIMENTAL PAGE

1. Place one drop of each of the indicated solutions on the acetate sheet. Record the color of each solution

NaCl	MgSO <sub>4</sub>	AlCl <sub>3</sub>
FeCl <sub>3</sub>	CaCl <sub>2</sub>	NiSO <sub>4</sub>
CuSO <sub>4</sub>	ZnCl <sub>2</sub>	AgNO <sub>3</sub>

2. A precipitate is a solid that separates upon mixing solutions. Predict which of the metal cations in this experiment will form colored precipitates upon the addition of NaOH. **That prediction is based on reading the introduction.** Add one drop of NaOH to find out. Record your results.

NaCl	MgSO <sub>4</sub>	AlCl <sub>3</sub>
FeCl <sub>3</sub>	CaCl <sub>2</sub>	NiSO <sub>4</sub>
CuSO <sub>4</sub>	ZnCl <sub>2</sub>	AgNO <sub>3</sub>

# Lab 10 Electron Configurations Of Atoms And Ions

**R Barnett**



## **Lab 10 Electron Configurations Of Atoms And Ions:**

## Whispering the Secrets of Language: An Psychological Quest through **Lab 10 Electron Configurations Of Atoms And Ions**

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