

Name: _____
Partner: _____

Date: _____

Types of Chemical Reactions

PRE-LAB:

There are many kinds of chemical reactions and several ways of classifying them. One popular and useful method is to classify reactions into four major types. They are: (1) direct combination, or synthesis; (2) decomposition; (3) single replacement; and (4) double replacement, also known as exchange of ions. Not all chemical reactions can be put into one of these categories. However, many can.

In a **synthesis** reaction, two or more substances combine to form a more complex substance. Equations for synthesis reactions have the general form $A + B \rightarrow AB$. An example of such a reaction is the formation of water from its constituent elements: $2H_2 + O_2 \rightarrow 2H_2O$.

A **decomposition** reaction is exactly the opposite of synthesis reaction. In a decomposition reaction, a compound breaks down into two or more simpler substances (elements or simpler compounds). The general form of the equation for a decomposition reaction is $AB \rightarrow A + B$. The breakdown of water into its elements is an example of such a reaction: $2H_2O \rightarrow 2H_2 + O_2$.

In a **single replacement** reaction, one substance in a compound is replaced by another, more active substance (element). Equations for single replacement reactions have two general forms. In reactions in which one metal replaces another metal, the general equation is $X + YB \rightarrow XB + Y$. In reactions in which one nonmetal replaces another nonmetal, the general form is $X + AY \rightarrow AX + Y$. The following equations illustrates the first of these types of single replacement reactions:

Zinc metal replaces copper (II) ion:



In a **double replacement** reaction, the metal ions of two different ionic compounds can be thought of as "replacing one another". Equations for this type of reaction have the general equation $AB + CD \rightarrow AD + CB$. Most "replacement" reactions (both single and double) take place in aqueous solutions containing free ions. In a double replacement, one of the products formed can be a precipitate, an insoluble gas, or water.

All of the types of reactions discussed here may be represented by balanced chemical equations. Reactions involving ion exchanges may also be represented by ionic equations. In this investigation you will be concerned only with chemical formulas and equations. In a balanced equation, the number of atoms of any given element must be the same on both sides of the equation. Multiplying the coefficient times the subscript of an element must yield the same product on both sides of the balanced equation.

In this experiment you will observe examples of the four types of chemical reactions described here. You will be expected to write and balance the equations representing the observed reactions.

PURPOSE:

Observe some chemical reactions and identify reactants and products of those reactions. Classify the reactions and write balanced equations.

EQUIPMENT:

Burner
Crucible tongs
Micro spatula
Test tubes

Test tube holder
Test tube rack
Wood splints
Sandpaper

Evaporating dish
Safety glasses

MATERIALS:

Zinc, mossy
Copper wire, 10cm
Magnesium ribbon, 5 cm
Copper (II) carbonate
3M Hydrochloric Acid

1M Copper (II) sulfate
0.1M Zinc acetate
0.1M Sodium phosphate
1M Sodium sulfite

Lab 14 Types Of Chemical Reactions Answers

James McKeen Cattell, Jaques Cattell



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Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001 *Catalog of Copyright Entries, Fourth Series* Library of Congress. Copyright Office, 1978 Energy Research Abstracts ,1979 **Nuclear Science Abstracts** ,1974 Differential Equations with Computer Lab Experiments Dennis G. Zill, 1995 Includes answers index **ERDA Energy Research Abstracts** United States. Energy Research and Development Administration. Technical Information Center, 1977 **Bibliography of Scientific and Industrial Reports** ,1965-07 *American Men of Science* James McKeen Cattell, Jaques Cattell, 1921 **Chemical Abstracts** ,1911 **Ceramic Abstracts** American Ceramic Society, 1926 **International Index to Periodicals** ,1924 **Decennial Index to Chemical Abstracts** ,1926 *Educational Film & Video Locator of the Consortium of College and University Media Centers and R.R. Bowker* Consortium of College and University Media Centers, 1990 *Practical Chemistry* Newton Henry Black, James Bryant Conant, 1920 *University of Nebraska-Lincoln, Catalog: ARTS & SCIENCES, COLLEGE OF*. University of Nebraska--Lincoln. College Of Arts & Sciences, 1928 **Industrial Arts Index** ,1923 Transactions of the American Nuclear Society American Nuclear Society, 1978 **Merck's Report** Theodore Weicker, 1902 **Bulletin of the Chemical Society of Japan** Nihon Kagakkai, 1982 An Introductory Guide to EC Competition Law and Practice Valentine Korah, 1994

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