

Annotated periodic table of the elements

<div><div>1</div><div>H1.008Hydrogen</div></div>																		<div><div>13</div><div>B10.81Boron</div></div>	<div><div>14</div><div>C12.01Carbon</div></div>	<div><div>15</div><div>N14.01Nitrogen</div></div>	<div><div>16</div><div>O16.00Oxygen</div></div>	<div><div>17</div><div>F18.99Fluorine</div></div>	<div><div>18</div><div>Ne20.18Neon</div></div>												
<div><div>3</div><div>Li6.94Lithium</div></div>		<div><div>4</div><div>Be9.012Beryllium</div></div>																	<div><div>19</div><div>Al26.98Aluminium</div></div>	<div><div>20</div><div>Si28.09Silicon</div></div>	<div><div>21</div><div>P30.97Phosphorus</div></div>	<div><div>22</div><div>S32.06Sulfur</div></div>	<div><div>23</div><div>Cl35.45Chlorine</div></div>	<div><div>24</div><div>Ar39.95Argon</div></div>											
<div><div>11</div><div>Na22.99Sodium</div></div>		<div><div>12</div><div>Mg24.31Magnesium</div></div>																	<div><div>25</div><div>Sc44.96Scandium</div></div>	<div><div>26</div><div>Ti47.87Titanium</div></div>	<div><div>27</div><div>V50.94Vanadium</div></div>	<div><div>28</div><div>Cr52.00Chromium</div></div>	<div><div>29</div><div>Mn54.94Manganese</div></div>	<div><div>30</div><div>Fe55.85Iron</div></div>	<div><div>31</div><div>Co58.93Cobalt</div></div>	<div><div>32</div><div>Ni58.69Nickel</div></div>	<div><div>33</div><div>Cu63.55Copper</div></div>	<div><div>34</div><div>Zn65.38Zinc</div></div>	<div><div>35</div><div>Ga69.72Gallium</div></div>	<div><div>36</div><div>Ge72.63Germanium</div></div>	<div><div>37</div><div>As74.92Arsenic</div></div>	<div><div>38</div><div>Se78.96Selenium</div></div>	<div><div>39</div><div>Br79.90Bromine</div></div>	<div><div>40</div><div>Kr83.80Krypton</div></div>	
<div><div>19</div><div>K39.10Potassium</div></div>		<div><div>20</div><div>Ca40.08Calcium</div></div>																	<div><div>37</div><div>Y88.91Yttrium</div></div>	<div><div>38</div><div>Zr91.22Zirconium</div></div>	<div><div>39</div><div>Nb92.91Niobium</div></div>	<div><div>40</div><div>Mo95.94Molybdenum</div></div>	<div><div>41</div><div>Tc98.91Technetium</div></div>	<div><div>42</div><div>Ru101.07Ruthenium</div></div>	<div><div>43</div><div>Rh101.07Rhodium</div></div>	<div><div>44</div><div>Pd106.32Palladium</div></div>	<div><div>45</div><div>Ag107.87Silver</div></div>	<div><div>46</div><div>Cd112.41Cadmium</div></div>	<div><div>47</div><div>In114.82Indium</div></div>	<div><div>48</div><div>Sn118.71Tin</div></div>	<div><div>49</div><div>Sb121.76Antimony</div></div>	<div><div>50</div><div>Te127.60Tellurium</div></div>	<div><div>51</div><div>I126.91Iodine</div></div>	<div><div>52</div><div>Xe131.29Xenon</div></div>	
<div><div>37</div><div>Rb85.47Rubidium</div></div>		<div><div>38</div><div>Sr87.62Strontium</div></div>																	<div><div>57 to 70</div><div></div></div>	<div><div>71</div><div>Lu174.97Lutetium</div></div>	<div><div>72</div><div>Hf178.49Hafnium</div></div>	<div><div>73</div><div>Ta180.95Tantalum</div></div>	<div><div>74</div><div>W183.84Tungsten</div></div>	<div><div>75</div><div>Re186.21Rhenium</div></div>	<div><div>76</div><div>Os190.23Osmium</div></div>	<div><div>77</div><div>Ir192.22Iridium</div></div>	<div><div>78</div><div>Pt195.08Platinum</div></div>	<div><div>79</div><div>Au196.97Gold</div></div>	<div><div>80</div><div>Hg200.59Mercury</div></div>	<div><div>81</div><div>Tl204.38Thallium</div></div>	<div><div>82</div><div>Pb207.2Lead</div></div>	<div><div>83</div><div>Bi208.98Bismuth</div></div>	<div><div>84</div><div>Po209Polonium</div></div>	<div><div>85</div><div>At210Astatine</div></div>	<div><div>86</div><div>Rn222Radon</div></div>
<div><div>55</div><div>Cs132.91Cesium</div></div>		<div><div>56</div><div>Ba137.33Barium</div></div>																	<div><div>89 to 102</div><div></div></div>	<div><div>103</div><div>La138.91Lanthanum</div></div>	<div><div>104</div><div>Rf181Rutherfordium</div></div>	<div><div>105</div><div>Db185Dubnium</div></div>	<div><div>106</div><div>Sg271Seaborgium</div></div>	<div><div>107</div><div>Bh270Bohrium</div></div>	<div><div>108</div><div>Hs277Hassium</div></div>	<div><div>109</div><div>Mt276Meitnerium</div></div>	<div><div>110</div><div>Ds285Darmstadtium</div></div>	<div><div>111</div><div>Rg281Roentgenium</div></div>	<div><div>112</div><div>Cn285Copernicium</div></div>	<div><div>113</div><div>Nh284Nihonium</div></div>	<div><div>114</div><div>Fl289Flerovium</div></div>	<div><div>115</div><div>Mc288Moscovium</div></div>	<div><div>116</div><div>Lv293Livermorium</div></div>	<div><div>117</div><div>Ts294Tennessine</div></div>	<div><div>118</div><div>Og294Oganesson</div></div>

Group names*

- Hydrogen & the alkali metals
- Alkaline earth metals
- Coinage metals
- Volatile metals
- Halogens
- Chalcogens
- Pnictogens
- Chalcogens
- Halogens
- Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
(Mulliken-Pauling)

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphoteric.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphoteric, and anionic species such as zincates, stannates, and bismuthates.

Metals

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

* as liquid P, the most stable form in ambient conditions.

Group 1

This group is Sc-Y, and the lanthanides (La-Lu) and the actinides (Ac-Lu). Trends going from Sc-Y and the lanthanides (La-Lu) trends going from Sc-Y and the actinides (Ac-Lu) trends going from Sc-Y and the actinides (Ac-Lu).

Nonmetals

Chalcogens
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal.

Lab 5 The Periodic Table Answers Key

RM Cervero



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