

# Specific Heat of a Metal Lab

## Objectives:

- Determine the specific heat of a metal.
- Calculate percent error of experimental results

## Introduction

Chemists identify substances on the basis of their chemical and physical properties. One physical property of a substance is the amount of energy it will absorb per unit of mass. This property can be measured quite accurately and is called *specific heat* ( $C_p$ ).

Specific heat is the amount of energy measured in joules, needed to raise the temperature of one gram of the substance one Celsius degree. Often applied to metallic elements, specific heat can be used as a basis for comparing energy absorption and transfer.

To measure specific heat in the laboratory a *calorimeter* of some kind must be used. A calorimeter is a well-insulated container used in measuring energy changes. The calorimeter is insulated to reduce the loss or gain of energy to or from the surroundings. Energy always flows from an object at a higher temperature to an object at a lower temperature. The heat gained by the cooler substance equals the heat lost by the warmer substance, if we assume no loss of heat to the surrounding environment.

$$\text{heat lost} = \text{heat gained}$$

In this experiment, you will determine the specific heat of a metal sample. The metal sample will be heated to a high temperature then placed into a calorimeter containing a known quantity of water at a lower temperature. Having measured the mass of the water in the calorimeter, the temperature change of the water ( $\Delta T$ ) and knowing the specific heat of water ( $4.184 \text{ J/g } ^\circ\text{C}$ ) the heat gained by the water (lost by the metal) can be calculated as follows:

$$Q = (m)(\Delta T)(C)$$

<i>heat gained by the water</i>	=	<i>mass of water (g)</i>	x	<i>change in temperature (<math>\Delta T</math>)</i>	x	<i>specific heat of water (<math>4.184 \text{ J/g } ^\circ\text{C}</math>)</i>
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*The specific heat of the metal can now be calculated:*

<i>Specific heat, <math>C</math> =</i>	$\frac{\text{heat gained by the water, } Q}{\text{of metal mass of metal (g) } m \times \Delta T \text{ of metal } (^\circ\text{C})}$
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**RC Schank**



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## **Table of Contents Lab 22 Specific Heat Of Metals Answers**

1. Understanding the eBook Lab 22 Specific Heat Of Metals Answers
  - The Rise of Digital Reading Lab 22 Specific Heat Of Metals Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Lab 22 Specific Heat Of Metals Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Lab 22 Specific Heat Of Metals Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 22 Specific Heat Of Metals Answers
  - Personalized Recommendations
  - Lab 22 Specific Heat Of Metals Answers User Reviews and Ratings
  - Lab 22 Specific Heat Of Metals Answers and Bestseller Lists
5. Accessing Lab 22 Specific Heat Of Metals Answers Free and Paid eBooks
  - Lab 22 Specific Heat Of Metals Answers Public Domain eBooks
  - Lab 22 Specific Heat Of Metals Answers eBook Subscription Services
  - Lab 22 Specific Heat Of Metals Answers Budget-Friendly Options

6. Navigating Lab 22 Specific Heat Of Metals Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Lab 22 Specific Heat Of Metals Answers Compatibility with Devices
  - Lab 22 Specific Heat Of Metals Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Lab 22 Specific Heat Of Metals Answers
  - Highlighting and Note-Taking Lab 22 Specific Heat Of Metals Answers
  - Interactive Elements Lab 22 Specific Heat Of Metals Answers
8. Staying Engaged with Lab 22 Specific Heat Of Metals Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Lab 22 Specific Heat Of Metals Answers
9. Balancing eBooks and Physical Books Lab 22 Specific Heat Of Metals Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Lab 22 Specific Heat Of Metals Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Lab 22 Specific Heat Of Metals Answers
  - Setting Reading Goals Lab 22 Specific Heat Of Metals Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab 22 Specific Heat Of Metals Answers
  - Fact-Checking eBook Content of Lab 22 Specific Heat Of Metals Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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