# 13.2

## THE NATURE OF LIQUIDS

## Section Review

#### Objectives

- Identify factors that determine physical properties of a liquid.
- Define avaporation in terms of kinetic energy
- Describe the equilibrium between a liquid and its vapor
- · Identify the conditions under which boiling occurs

#### Vocabulary

- vaporization
- evaporation.
- vapor pressure

- boiling point
- normal boiling point

## Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Liquids are much \_\_\_\_\_\_ than gases. Liquids and solids

are known as \_\_\_\_\_ states of matter. The conversion of a liquid

to a gas or vapor is called \_\_\_\_\_\_ When a liquid that is not
 \_\_\_\_\_\_ thanges to a gas, the process is called evaporation. A

Liquid evaporates faster when heated; however, evaporation itself

is a \_\_\_\_\_\_ process. When a partially filled container of liquid is

sealed, some of the particles at the \_\_\_\_\_\_ of the liquid vaporize.

These particles collide with the walls of the container, producing

a force called \_\_\_\_\_ The vapor pressure of a liquid can be

determined by a device called a \_\_\_\_\_ A liquid boils when its

\_\_\_\_\_\_ equals the external pressure. The normal boiling point of a

#### Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

 Particles in a liquid don't have enough kinetic energy to overcome the attractive forces between them and vaporize.

## **Nature Of Liquids Section Review Answers**

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lead to the following conclusion Water Science will continue to generate new material of utmost importance and high scientific quality for many more ISOPOW meetings to come

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