

## 2. Limiting and Excess Reactants

The **limiting reactant** is defined as the species in a chemical reaction that **theoretically** would be the first to be completely consumed if the reaction were to proceed to completion according to the chemical equation.

*- even if the reaction does not proceed to completion!*

All of the other reactant/s are called **excess reactants**.

For Example:



If 1 g mol of  $\text{C}_7\text{H}_{16}$  and 12 g mol of  $\text{O}_2$  are mixed so as to react,

# Limiting And Excess Reactants Pogil Key

**Lifelique**



**Limiting And Excess Reactants Pogil Key:**

*Limiting Reactant and Percent Yield*, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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