

# Limiting Reactant, Theoretical Yield, Actual Yield, & Percent Yield

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24



- If we have 53.2 g of Na and 65.8 g of  $\text{Cl}_2$ , what are the limiting reactant and theoretical yield?

$$53.2\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} \times \frac{2\text{ mol NaCl}}{2\text{ mol Na}} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 135.23$$

$$65.8\text{ g Cl}_2 \times \frac{1\text{ mol Cl}_2}{70.9\text{ g Cl}_2} \times \frac{2\text{ mol NaCl}}{1\text{ mol Cl}_2} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 110.5$$

# Limiting Reagent And Percent Yield Key

**Bill T Reese**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

### **Limiting Reagent And Percent Yield Key:**

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