

**MATHEMATICS GRADE 11
DEMARCATION
JUNE EXAMINATION 2024**

MATHEMATICS PAPER 1	MATHEMATICS PAPER 2
MARKS: 100 DURATION: 2 Hours Date: 31 May 2024	MARKS: 100 Duration: 2 Hours Date: 3 June 2024
CONTENT: <ul style="list-style-type: none"> • Algebra • Equations and inequalities • Exponents and Surds • Functions and Graphs 	CONTENT: <ul style="list-style-type: none"> • Analytical Geometry • Euclidean Geometry • Trigonometry

Mathematics Scope Paper1

LP Steffe



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Curriculum-based Assessment Carroll J. Jones, 2008 Although curriculum based assessment CBA has been on the horizon for a number of years there has been a need for a practical classroom based approach for its implementation This second edition is a major revision and update that offers practical and specific methods for developing and using CBAs in an educational setting It gives educators a highly specific step by step approach to building CBAs in the area of reading word recognition reading comprehension content reading mathematics and written expression Each chapter offers detailed easy to read and easy to follow instructions for the assessment construction process To further clarify the process extensive examples are given in table format A unique feature of the text is the manner in which the author illustrates the principles of CBA by providing a case study on one student which is referred to in each chapter Its formative assessment approach gives the teacher additional detailed information about students OCO performance which in turn should guide the type of instruction designed and implemented ultimately leading to higher performance on summative outcome measures Additionally detailed information is provided on subject area CBA construction the creation and implementation of a district wide CBA system for response to intervention and how to use IDEA S response to intervention in student evaluation Resources in Education ,2000-04 *The Student's Handbook to the University and Colleges of Cambridge* ,1913 **Proceedings of the National Academy of Sciences of the United States of America** National Academy of Sciences (U.S.),1915 The Proceedings of the National Academy of Sciences PNAS publishes research reports commentaries reviews colloquium papers and actions of the Academy PNAS is a multidisciplinary journal that covers the biological physical and social sciences **The Handbook - The Maharaja Sayajirao University of Baroda** Maharaja Sayajirao University of Baroda,1954 **Journal of Experimental Pedagogy and Training College Record** ,1923 The Forum of Education ,1923 *Forum of Education* ,1923 **Annals of Mathematics** ,1916 *Manitoba School Journal* ,1949 The Tammany Times ,1903 **Examination Decrees and Regulations** University of Oxford,1990 **Glasgow University Calendar** University of Glasgow,1909 **Glasgow University Calendar for the Year ...** University of Glasgow,1911 The Mathematics Educator ,2002 Calcutta Review ,1921 The Calendar University of Calcutta,1916 **Cambridge University Reporter** University of Cambridge,1966 Journal of the British Astronomical Association British Astronomical Association,1922 List of members 1890 1913 bound with v 1 23 **Parliamentary Papers** Great Britain. Parliament. House of Commons,1904

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was produced i what mass of iodine was produced

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h2s and multiply it with 2 mole h2o over 2 mole h2s from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

 8.3 mole h2s x 18g h2o i mole h2o 149 4g h2o is your theoretical yield so percent yield 137 1g h2o actual yield 149 4g h2o x 100
 91 77 problems 6 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o2 produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$$

 0 66 moles 32 g m

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained

calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160 0

g Al_2O_3 $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100 = 260.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of h 2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent

yield $\frac{\text{theoretical yield}}{\text{actual yield}} \times 100$ based on this definition we would expect a percent yield

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how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is

complete

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CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is

reacted with 8.6 g of H_2

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problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical

reaction

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 $80\text{ g I}_2\text{O}_5 \times \frac{1\text{ mol I}_2\text{O}_5}{283.3\text{ g I}_2\text{O}_5} \times \frac{1\text{ mol I}_2}{1\text{ mol I}_2\text{O}_5} \times \frac{253.8\text{ g I}_2}{1\text{ mol I}_2} = 70.9\text{ g I}_2$

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage
 $\text{percent yield} = \frac{\text{actual yield g}}{\text{theoretical yield g}} \times 100$ label 3 7 3