

Chapter 10 test**Matching**

Match each item with the correct statement below.

- | | |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition |
| b. mole | e. standard temperature and pressure |
| c. Avogadro's number | f. empirical formula |

- _____ 1. the number of representative particles of a substance present in 1 mole of that substance
_____ 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
_____ 3. the SI unit used to measure amount of substance
_____ 4. 0°C and 1 atm
_____ 5. the percent by mass of each element in a compound
_____ 6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 7. What SI unit is used to measure the number of representative particles in a substance?
a. kilogram
b. ampere
c. kelvin
d. mole
- _____ 8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C_3H_7O ?
a. $5 \times (6.02 \times 10^{23})$
b. 5
c. 35
d. $35 \times (6.02 \times 10^{23})$
- _____ 9. All of the following are equal to Avogadro's number EXCEPT _____.
a. the number of atoms of bromine in 1 mol Br_2
b. the number of atoms of gold in 1 mol Au
c. the number of molecules of nitrogen in 1 mol N_2
d. the number of molecules of carbon monoxide in 1 mol CO
- _____ 10. How many moles of tungsten atoms are in 4.8×10^{23} atoms of tungsten?
a. 8.0×10^{-7} moles
b. 8.0×10^{-1} moles
c. 1.3×10^{-4} moles
d. 1.3×10^{-2} moles
- _____ 11. How many atoms are in 0.075 mol of titanium?
a. 1.2×10^{23}
b. 2.2×10^{24}
c. 6.4×10^7
d. 4.5×10^{22}
- _____ 12. How many molecules are in 2.10 mol CO_2 ?
a. 2.53×10^{24} molecules
b. 3.79×10^{24} molecules
c. 3.49×10^{24} molecules
d. 1.26×10^{24} molecules

Modern Chemistry Chapter 10 Test

Fredus Nelson Peters



Modern Chemistry Chapter 10 Test:

Unveiling the Magic of Words: A Report on "**Modern Chemistry Chapter 10 Test**"

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