

# Review Questions

- What scientific law is explained using balanced chemical equations?
  - Law of chemical reactions
  - law of definite proportions
  - law of conservation of mass
  - Law of conservation of energy
- Which is not a type of chemical reaction?
  - Combination
  - decomposition
  - evaporation
  - Substitution
- Which statement best describes the purpose of balancing a chemical equation?
  - To satisfy the law of conservation of matter
  - To make the equation look more aesthetic
  - To create new elements and compounds
  - To make the equation harder to solve
- Why can't you change the formula in order to balance a reaction?
  - The molecule is indestructible.
  - The elements have definite mass ratios.
  - The number of atoms in the left should be equal to the atoms on the right of the arrow.
  - The elements that combine have definite composition and fixed proportion of elements by mass.
- Which of the following equations best represent the law of conservation of mass?
  - $1A + 2B \rightarrow 3AB$
  - $2A_2 + 8B \rightarrow 4AB_2$
  - $6A_2B \rightarrow 12A + B$
  - $2AB_3 \rightarrow 2A + 2B$
- When oxygen is available, sulfur dioxide is produced from the burning of sulfur. Which of the following word equations best represents this reaction?
  - sulfur + oxygen  $\rightarrow$  sulfur dioxide
  - sulfur dioxide + oxygen  $\rightarrow$  sulfur
  - sulfur dioxide  $\rightarrow$  sulfur + oxygen
  - sulfur  $\rightarrow$  sulfur dioxide + oxygen
- Which of the following is a balanced chemical equation?
  - $H_2O_2 \rightarrow H_2O + O_2$
  - $Na + Cl_2 \rightarrow NaCl$
  - $CO_2 + H_2O \rightarrow C_6H_{12}O_6 + O_2$
  - $AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$
- Which of these is an INCORRECT interpretation of this balanced equation?



- $2 \text{ g S} + 3 \text{ g O}_2 \rightarrow 2 \text{ g SO}_2$
  - $2 \text{ atoms S} + 3 \text{ molecules O}_2 \rightarrow 2 \text{ molecules SO}_2$
  - $2 \text{ moles S} + 3 \text{ moles O}_2 \rightarrow 2 \text{ moles SO}_2$
  - $64 \text{ g S} + 96 \text{ g O}_2 \rightarrow 2 \text{ g SO}_2 + 160 \text{ g SO}_2$
- In the reaction:  $Fe + O_2 \rightarrow Fe_2O_3$ , how many moles of  $Fe_2O_3$  are produced from 4 moles of Fe?
    - 1
    - 2
    - 3
    - 4
  - Magnesium reacts with oxygen gas to produce magnesium oxide. Which balanced chemical equation correctly represents this reaction?
    - $Mg + O_2 \rightarrow MgO$
    - $2Mg + O_2 \rightarrow 2MgO$
    - $Mg + O_2 \rightarrow 2MgO$
    - $2Mg + O_2 \rightarrow MgO$

# Modern Chemistry Review Questions

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