

Chapter 9: Stoichiometry

I. Stoichiometry

A. Relating the mass of the reactants with the mass of the products in a chemical reaction is Stoichiometry.

1. Stoichiometry follows the Law of Conservation of mass.
2. The total sum of the mass for reactants will equal the total sum of the mass of the products and NOT the number of molecules.
3. Stoichiometry is dependent on balanced equations.

B. Stoichiometric questions investigate the amount of substance needed to react with another substance, or the amount of a particular substance produced in a chemical reaction.

Stoichiometric relationships can look at:

1. mole to mole ratios.

a. Mole to mole ratios relate the number of moles of any substance in a balanced chemical reaction. They are based on the balanced chemical equation.



3 moles of H_2 requires 1 mole of N_2 to form 2 moles of NH_3 .

Therefore, 6 moles of H_2 should equal 4 moles of NH_3 .

This means that 17.25 moles of H_2 can produce 11.5 moles of NH_3 .

$$17.5 \text{ mol H}_2 \left(\frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) = 11.5 \text{ mol NH}_3$$

2. Mole to gram calculations.

a. Using a balanced chemical reaction the moles of one substance is given, and another substance has an equal amount of grams.

b. The equation MUST be balanced!



Using the same equation, how many grams of NH_3 are produced if 10.57 moles of H_2 are used in a reaction?

$$10.57 \text{ mol H}_2 \left(\frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} \right) \left(\frac{17.04 \text{ g NH}_3}{1 \text{ mol NH}_3} \right) = 120.08 \text{ g NH}_3$$

Mole to Mole ratio
from the BALANCED
chemical equation!

Mole to gram
conversion with the
molar mass from the
Periodic Table

d. The number of moles does NOT need to be a whole number.

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