



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant (K) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	 <p>Graphical Representation of Equilibrium</p> <p>The graph shows the concentration of reactants and products over time. The reactant concentration starts high and decreases exponentially towards a constant value. The product concentration starts at zero and increases exponentially towards a constant value. The equilibrium is reached when the concentrations of reactants and products remain constant over time.</p>
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant K_c is calculated using the concentrations of the reactants and products in moles per liter (M).</p> <p>For the reaction:</p> $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>The equilibrium constant K_c is given by:</p> $K_c = \frac{[NH_3]^2}{[N_2][H_2]^3}$	<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant K_p is calculated using the partial pressures of the reactants and products in atmospheres (atm).</p> <p>For the reaction:</p> $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>The equilibrium constant K_p is given by:</p> $K_p = \frac{(P_{NH_3})^2}{(P_{N_2})(P_{H_2})^3}$	<p>Equilibrium Constant, K_c vs. K_p</p> <p>The relationship between K_c and K_p is given by the equation:</p> $K_p = K_c(RT)^{\Delta n}$ <p>where R is the gas constant (0.0821 L atm / mol K) and Δn is the change in the number of moles of gas.</p>
<p>Le Chatelier's Principle</p> <p>Le Chatelier's Principle states that if a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift its equilibrium position to counteract the change.</p> <p>For example, if the concentration of a reactant is increased, the system will shift to the right to produce more products.</p>	<p>Effect of Concentration Changes</p> <p>Increasing the concentration of a reactant shifts the equilibrium to the right, favoring the forward reaction.</p> <p>Increasing the concentration of a product shifts the equilibrium to the left, favoring the reverse reaction.</p>	<p>Effect of Pressure Changes</p> <p>Increasing the pressure shifts the equilibrium to the side with fewer moles of gas.</p> <p>Decreasing the pressure shifts the equilibrium to the side with more moles of gas.</p>
<p>Effect of Temperature Changes</p> <p>For an exothermic reaction, increasing the temperature shifts the equilibrium to the left, favoring the reverse reaction.</p> <p>For an endothermic reaction, increasing the temperature shifts the equilibrium to the right, favoring the forward reaction.</p>	<p>Effect of Temperature Changes</p> <p>The equilibrium constant K is a function of temperature. For an exothermic reaction, K decreases as temperature increases. For an endothermic reaction, K increases as temperature increases.</p>	
3.4. Effect of Temperature Changes on the Equilibrium Constant		
<p>Exothermic</p>	<p>Equilibrium Constant</p>	<p>Temperature</p>
<p>Exothermic</p>	<ul style="list-style-type: none">Increasing the concentration of a reactant shifts the equilibrium to the right, favoring the forward reaction.Increasing the concentration of a product shifts the equilibrium to the left, favoring the reverse reaction.	<p>Increasing the temperature shifts the equilibrium to the left, favoring the reverse reaction.</p>
<p>Endothermic</p>	<ul style="list-style-type: none">Increasing the concentration of a reactant shifts the equilibrium to the right, favoring the forward reaction.Increasing the concentration of a product shifts the equilibrium to the left, favoring the reverse reaction.	<p>Increasing the temperature shifts the equilibrium to the right, favoring the forward reaction.</p>
<p>Exothermic</p>	<ul style="list-style-type: none">Increasing the concentration of a reactant shifts the equilibrium to the right, favoring the forward reaction.Increasing the concentration of a product shifts the equilibrium to the left, favoring the reverse reaction.	<p>Increasing the temperature shifts the equilibrium to the left, favoring the reverse reaction.</p>

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