

CHAPTER 9 REVIEW***Stoichiometry*****SECTION 1**

SHORT ANSWER Answer the following questions in the space provided.

1. _____ The coefficients in a chemical equation represent the
 - (a) masses in grams of all reactants and products.
 - (b) relative number of moles of reactants and products.
 - (c) number of atoms of each element in each compound in a reaction.
 - (d) number of valence electrons involved in a reaction.
2. _____ Which of the following would not be studied within the topic of stoichiometry?
 - (a) the mole ratio of Al to Cl in the compound aluminium chloride
 - (b) the mass of carbon produced when a known mass of sucrose decomposes
 - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
 - (d) the amount of energy required to break the ionic bonds in CaF_2
3. _____ A balanced chemical equation allows you to determine the
 - (a) mole ratio of any two substances in the reaction.
 - (b) energy released in the reaction.
 - (c) electron configuration of all elements in the reaction.
 - (d) reaction mechanism involved in the reaction.
4. _____ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
 - (a) reaction sequence.
 - (b) bond energy.
 - (c) mole ratio.
 - (d) element proportion.
5. Given the reaction represented by the following unbalanced equation: $\text{N}_2\text{O}(g) + \text{O}_2(g) \rightarrow \text{NO}_2(g)$
 - a. Balance the equation.

 - b. What is the mole ratio of NO_2 to O_2 ?

 - c. If 20.0 mol of NO_2 form, how many moles of O_2 must have been consumed?

 - d. Twice as many moles of NO_2 form as moles of N_2O are consumed. True or False?

 - e. Twice as many grams of NO_2 form as grams of N_2O are consumed. True or False?

Modern Chemistry Chapter 9 Mixed Review

Stoichiometry Answers

**Holt Rinehart & Winston, Holt,
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Introductory Chemistry Steven S. Zumdahl, 1996 *CHEMISTRY* SILBERBERG, 2003 **Chemical Engineering Catalog**, 1922 **Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001-01-01 **Modern Chemistry** Raymond E. Davis, 1999 2000 2005 State Textbook Adoption Rowan Salisbury **Formulation and Stoichiometry** Emil J. Margolis, 1968 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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