

HOLT MODERN CHEMISTRY – CHAPTER 1 REVIEW

The following pages contain the bulk (but not all) of the information for the chapter 1 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.
****** Test items will be recall, examples, and/or application of this content. ******

OUTCOMES

- 1.1: Distinguish between types of research – basic, applied, and technological development. (F & PK)
- 1.2: Distinguish between the physical and chemical properties of matter. (F & PK)
- 1.2: Distinguish among homogeneous and heterogeneous mixtures, elements, and compounds. (F & PK)
- 1.3: List properties of metals and non-metals. (T & R)
- 2.3: Determine significant figures. (F & PK)

1.1: CHEMISTRY IS A PHYSICAL SCIENCE

- **Vocabulary**
 - **chemistry** – the scientific study of the composition, structure, and properties of matter and the changes that matter undergoes
 - **chemical** – any substance with a defined composition
- **Chapter Highlights**
 - Chemistry is the study of the composition, structure, and properties of matter and the changes that matter undergoes.
 - A chemical is any substance that has a definite composition or is used or produced in a chemical process.
 - Basic research is carried out for the sake of increasing knowledge. Applied research is carried out to solve practical problems. Technological development involves the use of existing knowledge to make life easier or more convenient.

1.2: MATTER AND ITS PROPERTIES

- **Vocabulary**
 - **matter** – anything that has mass and takes up space
 - **mass** -- a measure of the amount of matter in an object
 - **element** – a substance that cannot be separated or broken down into simpler substances by chemical means; all atoms of an element have the same atomic number
 - **atom** – the smallest unit of an element that maintains the chemical properties of that element
 - **compound** -- a substance made up of atoms of two or more different elements joined by chemical bonds
 - **extensive property** -- a property that depends on the extent or size of a system
 - **intensive property** -- a property that does not depend on the amount of matter present, such as pressure, temperature, or density
 - **physical property** -- a characteristic of a substance that does not involve a chemical change, such as density, color, or hardness
 - **physical change** -- a change of matter from one form to another without a change in chemical properties
 - **change of state** -- the change of a substance from one physical state to another
 - **solid** -- the state of matter in which the volume and shape of a substance are fixed
 - **liquid** -- the state of matter that has a definite volume but not a definite shape
 - **gas** -- a form of matter that does not have a definite volume or shape
 - **plasma** -- in physical science, a state of matter that starts as a gas and then becomes ionized; it consists of freemoving ions and electrons and it takes on an electric charge
 - **chemical property** -- a property of matter that describes a substance's ability to participate in chemical reactions
 - **chemical change** -- a change that occurs when one or more substances change into entirely new substances with different properties
 - **chemical reaction** -- the process by which one or more substances change to produce one or more different substances
 - **reactant** -- a substance or molecule that participates in a chemical reaction
 - **product** -- a substance that forms in a chemical reaction

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Scott C. Dulebohn



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the thorniest and most pressing environmental questions that we face *New York Review of the Telegraph and Telephone and Electrical Journal* ,1901 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **The Saturday Review of Politics, Literature, Science and Art** ,1875 Environmental Chemistry Kenneth S. Overway, 2017-03-06 Covers the essentials of environmental chemistry and focuses on measurements that can be made in a typical undergraduate laboratory Provides a review of general chemistry nestled in the story of the Big Bang and the formation of the Earth Includes a primer on measurement statistics and quantitative methods to equip students to make measurements in lab Encapsulates environmental chemistry in three chapters on the atmosphere lithosphere and hydrosphere Describes many instruments and methods used to make common environmental measurements *The Electrical Review* ,1949 Boston Journal of Chemistry and Popular Science Review ,1871 **Iron & Coal Trades Review** ,1921 Vol 115 includes Diamond jubilee issue 1867 1927 Gas Journal ,1926 **Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity *Nature* Sir Norman Lockyer, 1911 **Nature** ,1875 *Modern Chemistry* Nicholas D. Tzimopoulos, 1993 **Metals and Alloys** ,1929 Introduction to Modern Chemistry: Student Manual Eugene Meyer, 1979 **Journal of Gas Lighting and Water Supply** ,1925

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