

Molarity and Dilution:

Molarity:

- When we talk about solutions, oftentimes instead of referring to a number of moles of a solute, we describe concentrations → this is how many moles of the solute are present per 1 litre of solution.
- The unit we use for this is molarity, which means moles of solute per litre of solution.
 - $M = \frac{\text{moles}}{\text{litre}}$
- For example, a 2 molar solution has 2 moles of solute for every litre of solution.
 - $2M = \frac{2 \text{ moles}}{\text{litre}}$
- If we want to know the molarity of a solution and we have the grams of solute, we just convert grams to moles using the molar mass of the substance, and divide by the total volume of the solution.
- Example: 0.100L H_2O solution with 10.0g NaCl:
 - $10.0g \text{ NaCl} \times \frac{1 \text{ mole NaCl}}{58.44g \text{ NaCl}} = 0.171 \text{ moles of NaCl}$
 - $\frac{0.171 \text{ mol NaCl}}{0.100L \text{ H}_2O} = 1.71 \text{ M NaCl solution}$

Dilution:

- We can alter the concentration of a solution by a process called dilution. This is the act of adding solvent to a solution. So if the moles of solute stay the same, the solution will become less concentrated as the volume increases.
- We make dilution calculations using this equation: $M_1V_1 = M_2V_2$
- Plug in the initial concentrations and volume of the solution, then the final volume you will dilute to, then solve for the new concentration.
 - $(1.0M)(1.0L) = M_2(2.0L)$
 - $\text{Molarity} = \frac{\text{moles of solute}}{\text{litres of solution}}$
 - $M_2 = \frac{1 \text{ mol}}{2 \text{ litres}} = 0.50 \text{ molarity (M)}$
- The quantities are inversely related, so doubling the volume will halve the concentration.

Practice problems:

What would be the molarity of a solution made from 86.5g $Ca(OH)_2$ in 200mL of water?

- $\text{Molarity} = \frac{\text{moles of solute}}{\text{litres of solution}}$
- $86.5g \text{ Ca(OH)}_2 \times \frac{1 \text{ mol Ca(OH)}_2}{74.1g \text{ Ca(OH)}_2} = 1.17 \text{ moles}$
- $M = \frac{1.17 \text{ moles}}{0.2L} = 5.85 \text{ M}$

To what volume would you need to dilute the above solution to make it 1.95 M?

- $M_1V_1 = M_2V_2$
- $(5.85M)(0.2L) = (1.95M) V_2$
- $\frac{5.85}{1.95} = 3$
- $0.2L \times 3 = 0.6L$
- $V_2 = 0.600L$

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Henry's Clinical Diagnosis and Management by Laboratory Methods John Bernard Henry, 2007 Rev ed of Clinical diagnosis and management by laboratory methods edited by John Bernard Henry 20th ed c2001

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Society of Automotive Engineers Handbook Society of Automotive Engineers, Incorporated, 1985 1983 SAE Handbook. Society of Automotive Engineers, 1983 Jena Review ,1980 **Annals of the New York Academy of Sciences** New York Academy of Sciences, 1971 Records of meetings 1808 1916 in v 11 27 *Energy Research Abstracts* ,1995 Semiannual with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works deriving from DOE other related government sponsored information and foreign nonnuclear information Arranged under 39 categories e g Biomedical sciences basic studies Biomedical sciences applied studies Health and safety and Fusion energy Entry gives bibliographical information and abstract Corporate author subject report number indexes *Water Research* ,1982 **Southern California Coastal Water Research Project Biennial Report** Southern California Coastal Water Research Project, 2001 *Mosby's Manual of Clinical Nursing Procedures* University of California, San Francisco. Department of Nursing Service, 1981 **Neonatal Enteric Infections Caused by Escherichia Coli** Bud Tennant, 1971 **Annual Report** Southern California Coastal Water Research Project, 2003 **Chemical Engineering Progress** ,1974 *Chemistry International* ,1992 *Biennial Report* Southern California Coastal Water Research Project, 2001

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