

## Mole Ratios

How can the coefficients in a chemical equation be interpreted?

NATALIA  
LACAYO  
P. 2

### Why?

A balanced chemical equation can tell us the number of reactant and product particles (ions, atoms, molecules or formula units) that are necessary to conserve mass during a chemical reaction. Typically when we balance the chemical equation we think in terms of individual particles. However, in real life the reaction represented by an equation occurs an unimaginable number of times. Short of writing very large numbers ( $10^{23}$  or larger) in front of each chemical in the equation, how can we interpret chemical equations so that they more realistically represent what is happening in real life? In this activity you will explore the different ways a chemical reaction can be interpreted.

### Model 1 – A Chemical Reaction



- Consider the reaction in Model 1.

a. What are the coefficients for each of the following substances in the reaction?

**1**  $\text{N}_2$

**3**  $\text{H}_2$

**2**  $\text{NH}_3$

b. Draw particle models below to illustrate the reaction in Model 1.



- Consider each situation below as it relates to the reaction in Model 1.

a. Calculate the amount of reactants consumed and products made.

b. Record the ratio of  $\text{N}_2$  to  $\text{H}_2$  to  $\text{NH}_3$ . Reduce the ratio to the lowest whole numbers possible.

|   | $\text{N}_2$<br>Consumed | $\text{H}_2$<br>Consumed | $\text{NH}_3$<br>Produced | Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$<br>(reduced) |
|---|--------------------------|--------------------------|---------------------------|--|
| For a single reaction, how many molecules of each substance would be consumed or produced?    | <b>1</b>                 | <b>3</b>                 | <b>2</b>                  | <b>1:3:2</b>   |
| If the reaction occurred one hundred times, how many molecules would be consumed or produced? | <b>100</b>               | <b>300</b>               | <b>200</b>                | <b>1:3:2</b>   |
| If the reaction occurred 538 times, how many molecules would be consumed or produced?         | <b>538</b>               | <b>1614</b>              | <b>1076</b>               | <b>1:3:2</b>   |

# Mole Ratio Pogil Key

**Lifelique**



**Mole Ratio Pogil Key:**

**Mole Ratios** Lifeliqe, 2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions

## The Enigmatic Realm of **Mole Ratio Pogil Key**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Mole Ratio Pogil Key** a literary masterpiece penned with a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of those that partake in its reading experience.

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