

Calculating Molar Mass

ANSWER KEY

Moles $\longleftrightarrow$ Grams

Directions: Find the percent composition of all the elements in each compound. Show all work. Percentages should be close to or exactly 100%. Round to two decimals.

Moles $\longrightarrow$ Grams

- How many grams of NaOH are in 1 mole? 40 g
 40×1
- How many grams of NaOH are in 0.1 mole? 4 g
 40×0.1
- What is the mass of 3 moles of O_2 ? 192 g
 32×3
- How many grams of $BaCl_2$ are in 0.25 moles? 52.1 g
 $208.25 \times .25$
- What is the mass of 4.1 moles of Fe_2O_3 ? 654.8 g
 159.7×4.1
- How many grams of $Ca(OH)_2$ are in 0.5 moles? 74.06 g
 148.1×0.5

Grams $\longrightarrow$ Moles

- How many moles are in 55.80 grams of Zn? 0.86
 $\frac{55.80}{64.90}$
- How many moles of KOH are there in 22.5 grams? 0.3
 $\frac{22.5}{96.1}$
- Determine the number of moles in 37.80 grams of $Ca(OH)_2$. 0.25
 $\frac{37.80}{152.98}$
- How many moles are in 9.5 g of F_2 ? 0.2
 $\frac{9.5}{48}$
- Convert 346.7 g of Al_2O_3 to moles 2.687
 $\frac{346.7}{129.08}$
- Find the number of moles in 46.7 grams of NaCl 0.8
 $\frac{46.7}{58.44}$

Name _____

Date _____

DR. DAVE'S
SCIENCE

CO_2

Element	Atoms	Atomic mass	Total mass
C	()	x	=
O	()	x	=

H_2SO_4

Element	Atoms	Atomic mass	Total mass
H	()	x	=
S	()	x	=
O	()	x	=

$PbCO_3$

Element	Atoms	Atomic mass	Total mass
Pb	()	x	=
C	()	x	=
O	()	x	=

$Mg(OH)_2$

Element	Atoms	Atomic mass	Total mass
Mg	()	x	=
O	()	x	=
H	()	x	=

$Al_2(SO_4)_3$

Element	Atoms	Atomic mass	Total mass
Al	()	x	=
S	()	x	=
O	()	x	=

Grams $\longrightarrow$ Moles
Moles $\longrightarrow$ Grams

Moles And Formulas Practice Skill Practice 4

Ying-Ying Zheng



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