

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

2. _____ According to the law of conservation of mass, the total mass of the reacting substances is
- always more than the total mass of the products.
 - always less than the total mass of the products.
 - sometimes more and sometimes less than the total mass of the products.
 - always equal to the total mass of the products.**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 2. In the reaction described by the word equation
- sodium + water \rightarrow sodium hydroxide + hydrogen gas
- sodium hydroxide is a product.
 - sodium is a product.
 - water is a reactant.
 - Both (a) and (c).**
- _____ 3. Which of the following indicates that a chemical equation is balanced?
- The numbers of atoms of each element are the same on both sides of the equation.**
 - All of the coefficients are the same.
 - The numbers of molecules on each side of the equation are equal.
 - The sums of the coefficients on each side of the equation are equal.
- _____ 4. From a complete and correctly written chemical equation, you can obtain the
- chemical formulas of the reactants and products.
 - relative amounts of the reactants and products.
 - physical states of the reactants and products.
 - All of the above**
- _____ 5. To balance a chemical equation, you must
- adjust the subscripts.
 - adjust the number of products.
 - adjust the number of reactants.
 - adjust the coefficients.**
- _____ 6. The formation of a solid product in a chemical reaction is represented by the symbol
- (aq).
 - (g).
 - (f).
 - (s).**
- _____ 8. In the expression 3CO_2 , the numbers 3 and 2 are, respectively,
- a subscript and a coefficient.
 - a coefficient and a subscript.**
 - two subscripts.
 - two coefficients.
- _____ 9. What is the total number of molecules that is represented by the following chemical equation? (Add up the coefficients.)
- $$\text{NH}_4\text{NO}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$$
- 3
 - 4**
 - 8
 - 16
- _____ 10. Which equation below violates the law of conservation of mass?
- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $\text{KCl} + \text{Br} \rightarrow \text{KBr} + \text{Cl}_2$**
 - $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
 - $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- _____ 11. Which of the following equations represents the balanced equation for the reaction of iron and oxygen?
- $2\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - $\text{Fe} + 3\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$**
 - $3\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- _____ 12. Which coefficients correctly balance the formula equation below?
- $$\text{KClO}_3(s) \rightarrow \text{KCl}(s) + \text{O}_2(g)$$
- 1, 1, 1
 - 1, 1, 3
 - 2, 2, 3**
 - 2, 1, 1
- _____ 13. Which equation is *not* balanced?
- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
 - $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$
 - $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$**
- _____ 14. In which type of reaction do two or more compounds react to form one product?
- synthesis (AKA combination)**
 - decomposition
 - combustion
 - single-replacement
- _____ 15. Which of the following is true regarding single-replacement reactions?
- Only one product is formed.
 - There is only one reagent.
 - One element replaces another in a compound.**

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Brendan G. Carr



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