

AP Chemistry Chapter 14 Answers – Zumdahl

14.27

a.
 $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightarrow \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a reaction as the equilibrium constant for the reaction is called K_a .

b.
 Propanoic acid is a weak acid, so it is only partially dissociated in water. The dissociation reaction is: $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+ + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$ or $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$.

c.
 NH_4^+ is a weak acid. Similar to propanoic acid, the dissociation reaction is:



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An acid is a proton (H^+) donor and a base is a proton acceptor. A conjugate acid-base pair differs by only a proton (H^+).

	Acid	Base	Conjugate Base of Acid	Conjugate Acid of Base
a.	H_2CO_3	H_2O	HCO_3^-	H_3O^+
b.	$\text{C}_5\text{H}_5\text{NH}^+$	H_2O	$\text{C}_5\text{H}_5\text{N}$	H_3O^+
c.	$\text{C}_6\text{H}_5\text{NH}^+$	HCO_3^-	$\text{C}_6\text{H}_5\text{N}$	H_2CO_3

14.31

Strong acids have a $K_a \gg 1$ and weak acids have $K_a < 1$. Table 14.2 in the text lists some K_a values for weak acids. K_a values for strong acids are hard to determine so they are not listed in the text. However, there are only a few common strong acids so if you memorize the strong acids, then all other acids will be weak acids. The strong acids to memorize are HCl , HBr , HI , HNO_3 , HClO_4 , and H_2SO_4 .

- HClO_4 is a strong acid.
- HOCl is a weak acid ($K_a = 3.5 \times 10^{-8}$).
- H_2SO_4 is a strong acid.
- H_2SO_3 is a weak diprotic acid with K_{a1} and K_{a2} values less than one.

14.33

The K_a value is directly related to acid strength. As K_a increases, acid strength increases. For water, use K_w when comparing the acid strength of water to other species. The K_a values are:



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Sabine Zange



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Modern Chemistry Nicholas D. Tzimopoulos,1993 **Electrochemistry in Nonaqueous Solutions** Kosuke

Izutsu,2009-09-22 An excellent resource for all graduate students and researchers using electrochemical techniques After introducing the reader to the fundamentals the book focuses on the latest developments in the techniques and applications in this field This second edition contains new material on environmentally friendly solvents such as room temperature ionic liquids Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001 *Principles of Modern*

Chemistry David W. Oxtoby,H. Pat Gillis,Laurie J. Butler,2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Foundations of College**

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Bockris,Amulya K.N. Reddy,2007-05-08 This book had its nucleus in some lectures given by one of us J O M B in a course on electrochemistry to students of energy conversion at the University of Pennsylv nia It was there that he met a number of people trained in chemistry physics biology metallurgy and materials science all of whom wanted to know something about electrochemistry The concept of writing a book about electrochemistry which could be understood by people with very varied backgrounds was thereby engendered The lectures were recorded and written up by Dr Klaus Muller as a 293 page manuscript At a later stage A K N R joined the effort it was decided to make a fresh start and to write a much more comprehensive text Of methods for direct energy conversion the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance Thus conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing

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V. Volckxsom,1882 English Mechanic and Mirror of Science ,1894 Modern Methods of Polymer Characterization
Howard G. Barth,Jimmy W. Mays,1991-09-03 Presents the methods used for characterization of polymers In addition to
theory and basic principles the instrumentation and apparatus necessary for methods used to study the kinetic and
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Castka,1966

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