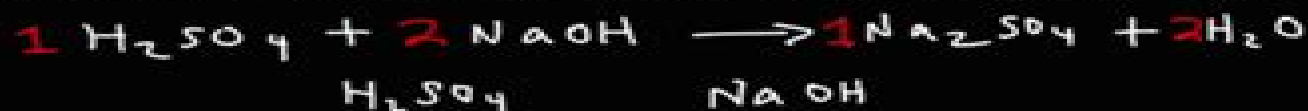


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1. 28.9 mL of H_2SO_4 was completely titrated with 38.4 mL of a 0.25M NaOH solution. What is the concentration of H_2SO_4 ?



$$2 M_1 V_1 = 1 M_2 V_2$$

$$2 M_1 (28.9 \text{ mL}) = 1 (0.25 \text{ M}) (38.4 \text{ mL})$$

$$\frac{\cancel{57.8} M_1}{\cancel{57.8}} = \frac{9.6}{57.8}$$

$$0.1661 \text{ M } \text{H}_2\text{SO}_4$$

$$M_1 = 0.1661 \text{ M}$$

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