



7. What is the maximum mass of aluminum hydroxide that can be produced when 25.0 mL of 0.248 M sodium hydroxide solution are mixed with 25.0 mL of 0.189 M aluminum nitrate solution?

Find mol of each reactant

$$0.0250 \text{ L} \times 0.248 \frac{\text{mol}}{\text{L}} = 6.2 \times 10^{-3} \text{ mol NaOH}$$

$$0.0250 \text{ L} \times 0.189 \frac{\text{mol}}{\text{L}} = 4.73 \times 10^{-3} \text{ mol Al}(\text{NO}_3)_3$$

$$6.20 \times 10^{-3} \text{ mol NaOH} \times \frac{1 \text{ mol Al}(\text{OH})_3}{3 \text{ mol NaOH}} \times \frac{78.0 \text{ g Al}(\text{OH})_3}{1 \text{ mol}} =$$

$$\boxed{0.161 \text{ g Al}(\text{OH})_3}$$

need 3 times as much NaOH as $\text{Al}(\text{NO}_3)_3$ but we only have 1.3 X as much NaOH is limiting

8. Indicate which of the following are redox reactions. For those that are redox reactions, label the reactant that is the oxidizing agent and the reactant that is the reducing agent. You do not have to balance these reactions!



Redox agent ox. agent

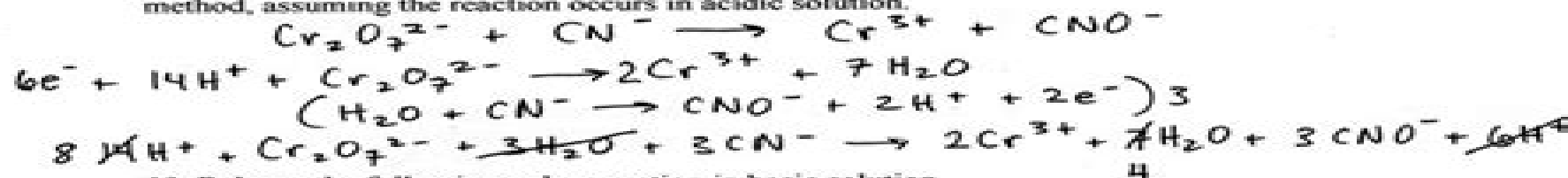


Not redox

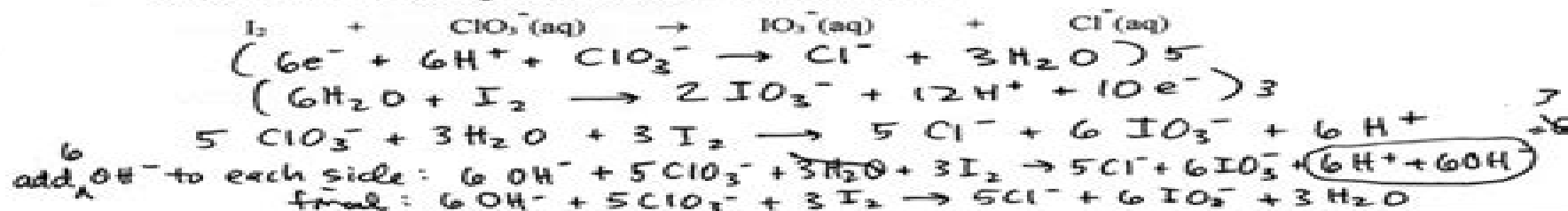


red agent ox agent

9. Aqueous dichromate ion reacts with aqueous cyanide ion (CN^-) to produce chromium (III) ion and aqueous cyanate (CNO^-) ion. Balance the equation using the half-reaction method, assuming the reaction occurs in acidic solution.



10. Balance the following redox reaction in basic solution.



Modern Chemistry Chapter 2 Homework

John L. Kice



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