

Q.3 The phenomenon in which an ionic compound shows more than one crystalline form is called:

- A) Polymorphism                      C) Isomorphism  
B) Allotropy                              D) Isomerism

Q.4 All of the following factors affect vapour pressure of liquid EXCEPT:

- A) Nature of liquid                      C) Surface area  
B) Temperature                        D) Intermolecular forces

Q.5 Which of the following gases shows more non-ideal behaviour?

- A)  $O_2$     C)  $N_2$   
B)  $CO_2$                                       D)  $H_2$

Q.6 Which of the following equations is used for real gases?

- A)  $PV = nRT$                                   C)  $\frac{PV}{T} = \frac{PV_0}{T_0}$   
B)  $PV = \frac{1}{3} n m C^2$                         D)  $\left(P + \frac{a}{V^2}\right)(V - b) = nRT$

Q.7 The gas laws can be summarised in the ideal gas equation  $PV = nRT$  which of the following statements is / are incorrect?

- A) One mole of any ideal gas occupies the same volume under the same conditions of temperature and pressure.  
B) The density of an ideal gas at constant pressure is inversely proportional to temperature.  
C) Volume of a given mass of a gas increases two times if temperature is raised from  $27^\circ\text{C}$  to  $54^\circ\text{C}$  at constant pressure.  
D) Both A and B

Q.8 Which of the following statements does not correctly represent the behaviour of an ideal gas?

- A)  $PV = C$                                       C)  $PM = nRT$   
B)  $P \propto CT$                                     D)  $P \propto \frac{1}{d}$

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**Jianjun Gao**



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## **Table of Contents Modern Chemistry Section 3 Gases Answer Key**

1. Understanding the eBook Modern Chemistry Section 3 Gases Answer Key
  - The Rise of Digital Reading Modern Chemistry Section 3 Gases Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Section 3 Gases Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Section 3 Gases Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Section 3 Gases Answer Key

- Personalized Recommendations
  - Modern Chemistry Section 3 Gases Answer Key User Reviews and Ratings
  - Modern Chemistry Section 3 Gases Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Section 3 Gases Answer Key Free and Paid eBooks
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    - Modern Chemistry Section 3 Gases Answer Key Budget-Friendly Options
  6. Navigating Modern Chemistry Section 3 Gases Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
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    - Modern Chemistry Section 3 Gases Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modern Chemistry Section 3 Gases Answer Key
    - Highlighting and Note-Taking Modern Chemistry Section 3 Gases Answer Key
    - Interactive Elements Modern Chemistry Section 3 Gases Answer Key
  8. Staying Engaged with Modern Chemistry Section 3 Gases Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modern Chemistry Section 3 Gases Answer Key
  9. Balancing eBooks and Physical Books Modern Chemistry Section 3 Gases Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modern Chemistry Section 3 Gases Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modern Chemistry Section 3 Gases Answer Key
    - Setting Reading Goals Modern Chemistry Section 3 Gases Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modern Chemistry Section 3 Gases Answer Key

- Fact-Checking eBook Content of Modern Chemistry Section 3 Gases Answer Key
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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