

CHEMICAL EQUILIBRIUM



EQUILIBRIUM CONSTANT 'K'.

For a General Reaction $aA + bB \rightleftharpoons pP + qQ$

The **equilibrium constant** expression is

$$K_c = \frac{[P]^p [Q]^q}{[A]^a [B]^b}$$

where K_c is the **Equilibrium Constant** (or K_p if they are all gases)

RELATION BETWEEN K_p AND K_c

For the Reaction $aA + bB \rightleftharpoons cC + dD$

$$K_p = \frac{[P_c]^c \times [P_D]^d}{[P_A]^a \times [P_B]^b} = \frac{[C]^c [D]^d}{[A]^a [B]^b} \frac{(RT)^{c+d}}{(RT)^{a+b}}$$

$$K_p = K_c (RT)^{\Delta n_g}$$

$$\text{if } \Delta n_g = 0 \implies K_p = K_c$$

Where, $\Delta n_g = (c+d) - (a+b)$

= no. of moles of gaseous products - no. of moles of gaseous Reactants

WHAT DOES THE VALUE OF 'K' MEAN ?

(a) $K \gg 1$

Reactants



Products

If $K \gg 1$, the reaction is product-favoured; product predominates at **Equilibrium**.

(a) $K \ll 1$

Reactants



Products

If $K \ll 1$, the reaction is reactant-favoured; reactant predominates at **Equilibrium**.

(a) $K = 1$

The reaction lies in the middle (mix of reactants and products)

MAGNITUDE OF 'K'

Small ($K < 10^{-3}$)



Mostly Reactants

Intermediate ($10^{-3} \leq K \leq 10^3$)



Significant amounts of reactants and products

Large ($K > 10^3$)



Mostly Products

Modern Chemistry Section Chemical Equilibrium

J Ma



Modern Chemistry Section Chemical Equilibrium:

Whispering the Secrets of Language: An Psychological Quest through **Modern Chemistry Section Chemical Equilibrium**

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