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# > Coursebook answers

## Chapter 16



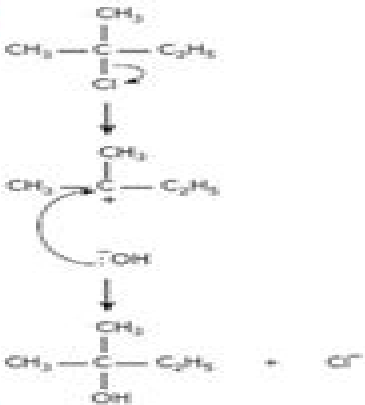
### Science in context

Discuss as a group how the information about the use of halogenoalkanes demonstrates how some scientific developments can have both good and bad effects on society and the environment.

- Having read the Science in Context passage, ask learners to shut their Coursebook. Then conduct a quick activity asking individual learners to give a specific example of a beneficial use of halogenoalkanes or a problematic use, until the class runs out of suggestions.
- You might then ask for a class vote, show of hands, on whether the halogenoalkanes have benefitted or hindered human progress in the last century.

### Self-assessment questions

- propene and chlorine
  - $\text{OH}^-$  ions are negatively charged so are attracted more strongly than neutral water molecules to the partially positively charged carbon atoms in halogenoalkanes.
  - The hydrolysis of halogenoalkanes produces halide ions; the rate of their formation can be monitored by using silver nitrate solution. The silver halide precipitates make the reaction mixture cloudy. The ionic equations for the formation of the precipitates are:  
 $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$   
 $\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$   
 $\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
  - Both ammonia and amines contain a nitrogen atom with a lone pair of electrons that is available to donate.
  - tripropylamine  $(\text{CH}_3\text{CH}_2\text{CH}_2)_3\text{N}$

- 
  - 
  - 
- $\text{CH}_3\text{CHBrCH}_3 + \text{NaOH}(\text{ethanol}) \rightarrow \text{CH}_2=\text{CHCH}_3 + \text{H}_2\text{O} + \text{NaBr}$
  - propene
  - $\text{BrCH}_2\text{CH}_2\text{CH}_3 + \text{NaOH}(\text{ethanol}) \rightarrow \text{CH}_2=\text{CHCH}_3 + \text{H}_2\text{O} + \text{NaBr}$   
 (1-bromopropane)

# Modern Chemistry Text Answers Chapter 16

**John Krenos, Joseph Potenza, Loretta  
Jones, Lynn Koplitz, Thomas Spence**



## **Modern Chemistry Text Answers Chapter 16:**

*Modern Chemistry* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001      *Chemical Principles Student's Study Guide & Solutions Manual* John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text      **Study and**

**Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb** Wade A. Freeman, 1986

Introduction to Modern Chemistry Edward Florian Neuzil, 1968      *Molecular Driving Forces* Ken Dill, Sarina Bromberg, 2010-10-21 *Molecular Driving Forces* Second Edition E book is an introductory statistical thermodynamics text that describes the principles and forces that drive chemical and biological processes It demonstrates how the complex behaviors of molecules can result from a few simple physical processes and how simple models provide surprisingly accurate insights into the workings of the molecular world Widely adopted in its First Edition *Molecular Driving Forces* is regarded by teachers and students as an accessible textbook that illuminates underlying principles and concepts The Second Edition includes two brand new chapters 1 Microscopic Dynamics introduces single molecule experiments and 2 Molecular Machines considers how nanoscale machines and engines work The Logic of Thermodynamics has been expanded to its own chapter and now covers heat work processes pathways and cycles New practical applications examples and end of chapter questions are integrated throughout the revised and updated text exploring topics in biology environmental and energy science and nanotechnology Written in a clear and reader friendly style the book provides an excellent introduction to the subject for novices while remaining a valuable resource for experts      **Chemical Principles Study Guide/Solutions Manual** John

Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses *Chemical Principles* helps students develop chemical insight by showing the connection between chemical principles and their applications      **Chemistry**

Leonard W. Fine, 1972      **Chemistry** Theodore Lawrence Brown, 1988      **Materials Science for Engineers** J.C.

Anderson, Keith D. Leaver, Rees D. Rawlings, Patrick S. Leever, 2004-06-01 This fifth edition of a successful textbook continues to provide students with an introduction to the basic principles of materials science over a broad range of topics The authors have revised and updated this edition to include many new applications and recently developed materials The book is presented in three parts The first section discusses the physics chemistry and internal structure of materials The second part examines the mechanical properties of materials and their application in engineering situations The final section presents the electromagnetic properties of materials and their application Each chapter begins with an outline of the relevance of its topics and ends with problems that require an understanding of the theory and some reasoning ability to resolve These are followed by self assessment questions which test students understanding of the principles of materials science and are designed to quickly cover the subject area of the chapter This edition of *Materials Science for Engineers*

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