

Key:

## Molarity Practice Problems

Show your work and place your answer with proper significant figures.

1. When the molarity of 150. mL of an aqueous calcium chloride solution is 0.895M, how many grams of the solute are present?

$$\frac{0.895 \text{ moles}}{\text{L}} \times \frac{0.150 \text{ L}}{1} = \frac{110.98 \text{ g CaCl}_2}{1 \text{ mole CaCl}_2} = 14.9 \text{ g CaCl}_2$$

2. What is the molarity of a solution made by adding 0.500g of sodium hydroxide to enough water to make 200. mL of solution?

$$0.500 \text{ g NaOH} \times \frac{1 \text{ mole NaOH}}{39.997 \text{ g NaOH}} = \frac{1}{0.200 \text{ L}} = 0.0625 \text{ M}$$

3. If the molarity of an aqueous solution of aluminum chloride is 0.270M, how many grams of aluminum chloride are present in 1.50L of water?

$$\frac{0.270 \text{ moles AlCl}_3}{\text{Liter}} \times \frac{1.50 \text{ L}}{1} = \frac{133.34 \text{ g AlCl}_3}{1 \text{ mole AlCl}_3} = 54.0 \text{ g AlCl}_3$$

4. What is the molarity of 250. mL of an aqueous solution containing 10.0g of potassium dichromate?

$$10.0 \text{ g K}_2\text{Cr}_2\text{O}_7 \times \frac{1 \text{ mole K}_2\text{Cr}_2\text{O}_7}{294.185 \text{ g K}_2\text{Cr}_2\text{O}_7} = \frac{1}{0.250 \text{ L}} = 0.136 \text{ M}$$

5. How many grams of sodium hydroxide are needed to make 100. mL of a 2.5M solution of sodium hydroxide?

$$\frac{2.5 \text{ moles}}{\text{L}} \times \frac{0.100 \text{ L}}{1} = \frac{39.997 \text{ g NaOH}}{1 \text{ mole NaOH}} = 10. \text{ g NaOH}$$

# Molarity Practice Answers With Work

**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**



## **Molarity Practice Answers With Work:**

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