

## Mole Ratios

How can the coefficients in a chemical equation be interpreted?

### Why?

A balanced chemical equation can tell us the number of reactant and product particles (ions, atoms, molecules or formula units) that are necessary to conserve mass during a chemical reaction. Typically when we balance the chemical equation we think in terms of individual particles. However, in real life the reaction represented by an equation occurs an unimaginable number of times. Short of writing very large numbers ( $10^{23}$  or larger) in front of each chemical in the equation, how can we interpret chemical equations so that they more realistically represent what is happening in real life? In this activity you will explore the different ways a chemical reaction can be interpreted.

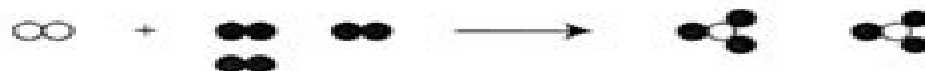
### Model 1 – A Chemical Reaction



1. Consider the reaction in Model 1.
  - a. What are the coefficients for each of the following substances in the reaction?



- b. Draw particle models below to illustrate the reaction in Model 1.



2. Consider each situation below as it relates to the reaction in Model 1.
  - a. Calculate the amount of reactants consumed and products made.
  - b. Record the ratio of  $\text{N}_2$  to  $\text{H}_2$  to  $\text{NH}_3$ . Reduce the ratio to the lowest whole numbers possible.

	$\text{N}_2$ Consumed	$\text{H}_2$ Consumed	$\text{NH}_3$ Produced	Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$ (reduced)
For a single reaction, how many molecules of each substance would be consumed or produced?	1 molecule	3 molecules	2 molecules	1:3:2
If the reaction occurred one hundred times, how many molecules would be consumed or produced?	100 molecules	300 molecules	200 molecules	1:3:2
If the reaction occurred 538 times, how many molecules would be consumed or produced?	538 molecules	1614 molecules	1076 molecules	1:3:2

# Mole Ratios Pogil Answers

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## **Table of Contents Mole Ratios Pogil Answers**

1. Understanding the eBook Mole Ratios Pogil Answers
  - The Rise of Digital Reading Mole Ratios Pogil Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mole Ratios Pogil Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mole Ratios Pogil Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Ratios Pogil Answers
  - Personalized Recommendations
  - Mole Ratios Pogil Answers User Reviews and Ratings
  - Mole Ratios Pogil Answers and Bestseller Lists
5. Accessing Mole Ratios Pogil Answers Free and Paid eBooks
  - Mole Ratios Pogil Answers Public Domain eBooks
  - Mole Ratios Pogil Answers eBook Subscription Services
  - Mole Ratios Pogil Answers Budget-Friendly Options
6. Navigating Mole Ratios Pogil Answers eBook Formats

- ePub, PDF, MOBI, and More
- Mole Ratios Pogil Answers Compatibility with Devices
- Mole Ratios Pogil Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mole Ratios Pogil Answers
  - Highlighting and Note-Taking Mole Ratios Pogil Answers
  - Interactive Elements Mole Ratios Pogil Answers
- 8. Staying Engaged with Mole Ratios Pogil Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mole Ratios Pogil Answers
- 9. Balancing eBooks and Physical Books Mole Ratios Pogil Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mole Ratios Pogil Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mole Ratios Pogil Answers
  - Setting Reading Goals Mole Ratios Pogil Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mole Ratios Pogil Answers
  - Fact-Checking eBook Content of Mole Ratios Pogil Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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