

Molecular Composition of Gases

CHECK YOUR KNOWLEDGE

QUESTION 1 Answer the following questions in the space provided.

1. _____ The average speed of a gas molecule is most directly related to the _____.
 - (a) volume of the container
 - (b) pressure within the gas
 - (c) temperature of the gas
 - (d) number of molecules in the gas
2. _____ Which value below best represents the average molecular speed of a diatomic gas given at the same temperature?
 - (a) The value associated with molecular mass is directly proportional.
 - (b) The value associated with molecular mass is directly proportional.
 - (c) The value associated with square root of molecular mass is directly proportional.
 - (d) The value associated with the square of molecular mass is directly proportional.
3. Which combination of conditions corresponds to the following equation?

$$\text{volume} \propto \text{pressure} \propto \text{temperature}$$

 - a. _____ $\text{PV} = nRT$ is correct. The volume of H_2 is constant and the value of H_2 is constant.
 - b. _____ $\text{PV} = nRT$ is correct. The value of H_2 is constant and the value of H_2 is constant.
4. Which of the following is not a property of an ideal gas?
 - a. _____ The mass of the gas is proportional to the volume of the container.
 - b. _____ The mass of the gas is proportional to the volume of the container.
 - c. _____ The mass of the gas is proportional to the volume of the container.

QUESTION 2 The gas that occupies the most space is the one that has the least mass. Which of the following is the gas that occupies the most space?

1. _____ Which of the following is the gas that occupies the most space?

Molecular Composition Of Gases Mixed Review

Michael Hoelscher



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