$$\mathbf{r}_{A2B} = \frac{\mathbf{K}_3 \frac{\mathbf{K}_1 [A]^2}{2} [B]}{[\mathbf{K}_3 [B] + \mathbf{K}_2]}$$

If K_2 is very small, $K_2 \cong 0$

$$\mathbf{r}_{A2B} = \frac{\mathbf{K}_1 \mathbf{K}_3 [\mathbf{A}]^2 [\mathbf{B}]}{\mathbf{K}_3 [\mathbf{B}]}$$

So, from above equation we can conclude that our first assumption is wrong so, mechanism also wrong.

Now, we can assume for second order reaction

$$A + B \leftrightarrow AB^*$$
 $AB^* + A \leftrightarrow A_2B$
 $AB^* + A \leftrightarrow A_2B$
 $AB^* + AB^* + AB^*$
 $AB^* + AB^*$
 AB^*

From reaction (3) and (4) rate of formation of r[A2B],

$$\mathbf{r}_{AB^*} = \mathbf{K}_1[A][B] - \mathbf{K}_2[AB^*] - \mathbf{K}_3[AB^*][B] + \mathbf{K}_4[A_2B] \quad ..(5)$$

$$\mathbf{r}[AB^*] = \mathbf{0}$$

$$\mathbf{0} = \mathbf{K}_1[A][B] - \mathbf{K}_2[AB^*] - \mathbf{K}_3[AB^*][B] + \mathbf{K}_4[A_2B]$$

$$[AB^*] = \frac{\mathbf{K}_1[A][B] + \mathbf{K}_4[A_2B]}{[\mathbf{K}_3[A] + \mathbf{K}_2]}$$

Put [AB"] value in equation (5),

$$r_{A2B} = K_3 \left(\frac{K_1[A][B] + K_4[A_2B]}{[K_3[A] + K_2]} \right) [A] - K_4[A_2B]$$

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Paul Ih-fei Liu

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- 1. Understanding the eBook Levenspiel Solution Manual
 - The Rise of Digital Reading Levenspiel Solution Manual
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Levenspiel Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Levenspiel Solution Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Levenspiel Solution Manual
 - Personalized Recommendations
 - Levenspiel Solution Manual User Reviews and Ratings
 - $\circ\,$ Levenspiel Solution Manual and Bestseller Lists
- 5. Accessing Levenspiel Solution Manual Free and Paid eBooks
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 - Levenspiel Solution Manual eBook Subscription Services
 - Levenspiel Solution Manual Budget-Friendly Options
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- Levenspiel Solution Manual Compatibility with Devices
- Levenspiel Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Levenspiel Solution Manual
 - Highlighting and Note-Taking Levenspiel Solution Manual
 - Interactive Elements Levenspiel Solution Manual
- 8. Staying Engaged with Levenspiel Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - $\circ\,$ Following Authors and Publishers Levenspiel Solution Manual
- 9. Balancing eBooks and Physical Books Levenspiel Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Levenspiel Solution Manual
- 10. Overcoming Reading Challenges
 - $\circ\,$ Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Levenspiel Solution Manual
 - Setting Reading Goals Levenspiel Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Levenspiel Solution Manual
 - Fact-Checking eBook Content of Levenspiel Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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