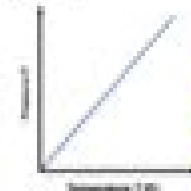
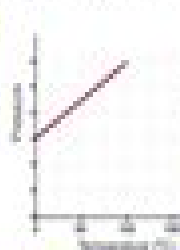
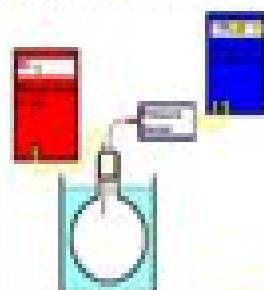


## GAS LAWS SUMMARY

## 1. Pressure and Temperature

(Volume constant, mass constant)



Pressure is directly proportional to Temperature

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

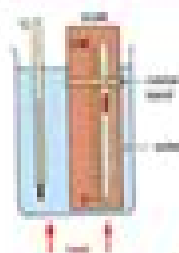
## Explanatory

Temperature ↑ particles have greater speed, and greater KE

 Particles collide with container walls more often and with greater force → Pressure ↑ (since  $P = F/A$ , Pressure ↑)

## 2. Volume and Temperature

(Pressure constant, mass constant)



Volume is directly proportional to Temperature

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

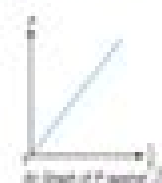
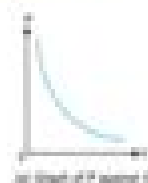
## Explanatory

Volume ↑ particles collide with container walls more often

 (Small force increases → since  $P = F/A$ , Pressure ↑)

## 3. Pressure and Volume

(Temperature constant, mass constant)



$$P_1 V_1 = P_2 V_2$$

## Explanatory

Volume ↓ particles collide with container walls more often

 (Small force increases → since  $P = F/A$ , Pressure ↑)

Pressure is the force per unit area

$$P = \frac{F}{A}$$

 $P = \text{the pressure (N/m}^2\text{)} = \frac{F}{A} = \frac{\text{N}}{\text{m}^2}$ 

Combined Gas Equation

$$\frac{P}{T} = \text{constant}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

(degrees Celsius or kelvin × 273)

(kelvin to degrees Celsius - 273)

Absolute Zero: 273 K / 0 °C

 There is no negative on the kelvin scale  
Temperature is a measure of the mean kinetic energy of particles

SPECIFIC HEAT CAPACITY (c): Heat energy required to raise the temperature of 1 kg of a substance by 1°C

$$E_h = c m \Delta T$$

 $E_h$ : heat energy (J)

 $c$ : specific heat capacity (J kg<sup>-1</sup> °C<sup>-1</sup>)

 $m$ : mass (kg)

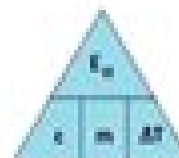
 $\Delta T$ : change in Temperature (°C)

 Specific heat capacity values are different for all materials  
- values found in data sheet

- values found in data sheet

 $E = P t$  can be used to find heat energy supplied by appliances

Heat energy will usually be lost to the surroundings - not all will be transferred to the substance being heated



Low specific heat capacity -

not much energy to heat (heats up quickly for constant power) but loses heat quickly

High specific heat capacity -

lots of energy to heat (heats up slowly for constant power) but loses heat slowly (retains heat better)

Specific Latent Heat (L): Heat energy required to change the state of 1 kg of a substance at the same temperature

$$E_h = m L$$

 $E_h$ : heat energy (J)

 $m$ : mass (kg)

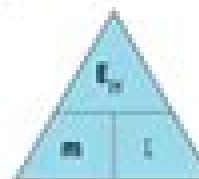
Specific latent heat values are different for all materials - values found in data sheet

 Specific Latent heat of fusion,  $L_f$ :

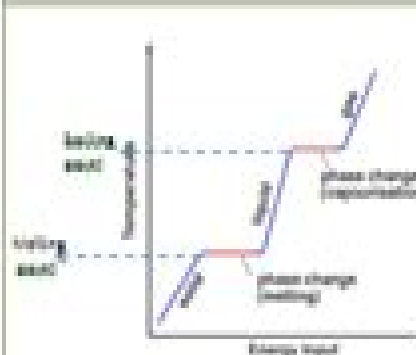
Solid → Liquid

 Specific Latent heat of vaporisation,  $L_v$ :

Liquid → Gas


 $L$ : specific latent heat of fusion / vaporisation

The same material requires different quantities of heat to change the state of unit mass from solid to liquid (fusion) and to change the state of unit mass from liquid to gas (vaporisation)



During change of state

- temperature remains constant

For change of temperature:

$$E_h = c m \Delta T$$

For change of state:

$$E_h = m L$$

"specific heat" = per kg diagram

# Memorandum Of Engineering Physics N5 For November 2013

**Tobias Bleicker**



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