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ADDITIONAL MATERIALS

The use of a calculator is not permitted in this examination.
A ruler, a protractor and a pair of compasses may be required.

INSTRUCTIONS TO CANDIDATES

Use black ink or black ball-point pen. Do not use gel pen or correction fluid.

You may use a pencil for graphs and diagrams only.

Write your name, centre number and candidate number in the spaces at the top of this page.

Answer all the questions in the spaces provided.

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Take a rest 15 minutes.

INFORMATION

You should give details of your method of solution when appropriate.

Answers without diagrams are not shown to scale.

Scale drawing solutions will not be awarded when you are asked to calculate.

The number of marks is given in brackets at the end of each question or part question.

In question 18, the assessment will take into account the quality of your algebraic and mathematical organisation, communication and accuracy in writing.

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GCSE

MATHEMATICS UNIT 1 – NON-CALCULATOR FOUNDATION TIER

1 hour 30 minutes

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Question	Maximum Mark	Mark Awarded
1.	5	
2.	5	
3.	5	
4.	5	
5.	5	
6.	5	
7.	5	
8.	5	
9.	5	
10.	5	
11.	5	
12.	5	
13.	5	
14.	5	
15.	5	
16.	5	
17.	5	
18.	5	
19.	5	
20.	5	
21.	5	
22.	5	
23.	5	
24.	5	
25.	5	
26.	5	
27.	5	
28.	5	
29.	5	
30.	5	
Total	150	

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Wyn Brice,Linda Mason,Tony Timbrell



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web one of the ions produced is always h which immediately combines with a water molecule to form the hydronium ion h₃o⁺ the h₃o⁺ ion is what defines the acidic properties of a substance because of their special classification acids have a naming system different from ionic or other molecular covalent compounds

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brf h hhhhhf hcl hbr hhfcl hbr hfformulahcl hbr hfnamehydrochloric acid

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web the molecular or empirical formula acids have their own set of naming rules cf sec 2 8 which you must also know you also need to be able to recognize and name a few organic compounds such as simple alkanes and alcohols see sec 2 9 key questions 8 in the chemical formula and name which element is given first a metal or nonmetal 9

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web each pair has one acid that ends in ic and another that ends in ous these endings are related to the name of the polyatomic anion found in the acid ate or ite complete the statements below with the correct acid name ending polyatomic anion ending is ate acid name ending is if