

25. The number of  $\sigma$  bonds in  $N \equiv N$  is
- 1
  - 2
  - 3
  - 4
26. The elements in an ionic compound are held together by
- electrostatic forces of attraction.
  - van der Waals forces
  - the spin of paired electrons.
  - the formation of hybrid orbitals.
  - an electron pair.
27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
- at which oxidation occurs.
  - which attracts cations.
  - at which electrons are supplied to the solution.
  - at which reduction occurs.
28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
- 47
  - 62
  - 93
  - 186
29. In a galvanic (voltaic) cell in which the reaction is  $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$  and the ions are at unit concentration (activity), the cell potential is
- $$\begin{array}{ll} Cd \rightarrow Cd^{2+} + 2e^- & 0.4021 \text{ V} \\ Cu \rightarrow Cu^{2+} + 2e^- & -0.344 \text{ V} \end{array}$$
- 0.1383 V
  - 0.4021 V
  - 0.344 V
  - 0.7461 V
  - 0.3677 V
30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
- $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$
  - $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
  - $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$
  - $COCl_2(g) \rightleftharpoons CO(g) + Cl_2(g)$

Standard Potentials	$E^\circ$
$Mg \rightarrow Mg^{2+} + 2e^-$	2.37V
$Al \rightarrow Al^{3+} + 3e^-$	1.66V
$Zn \rightarrow Zn^{2+} + 2e^-$	0.76V
$Fe \rightarrow Fe^{2+} + 2e^-$	0.44V
$Cu \rightarrow Cu^{2+} + 2e^-$	0.34V
$Ag \rightarrow Ag^+ + e^-$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
- Mg and Ag
  - Mg and Fe
  - Zn and Cu
  - Al and Ag
  - Mg and Al
32.  $E = E^\circ - 0.059/n \log Q$  (Nernst equation)  
 $[H^+] = 1.0 \text{ M}$  initially,  $P_{O_2} = 1.0 \text{ atm}$
- $$4e^- + O_2(g) + 4H^+(aq) \rightleftharpoons 2H_2O(l) \quad E^\circ = 1.23 \text{ V}$$
- Based on the information above, which statement is correct?
- $n = 1$ , since one mole of oxygen is being considered.
  - Addition of base should result in an  $E$  value, which is less than 1.23 V.
  - $E$  is independent of the pH of the solution.
  - $Q = \frac{[H_2O]^2}{[O_2][H^+]^4}$
33. The equilibrium constant for the gaseous reaction  $C + D \rightleftharpoons E + 2F$  is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
- proceed at equal rates in both directions.
  - proceed more rapidly to form E and F.
  - proceed more rapidly to form C and D.
  - not occur in either direction.
- | Compound  | $\Delta G_f^\circ$ , kJ/mol |
|-----------|-----------------------------|
| $H_2O(l)$ | -237                        |
| $H_2O(g)$ | -229                        |
34. At 298 K the equilibrium constant for  $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(l)$
- is larger than the  $K_{eq}$  for  $H_2(g) + \frac{1}{2} O_2(g) \rightleftharpoons H_2O(g)$
  - will have a value of 1.0 at equilibrium.
  - cannot be computed since data on  $O_2$  and

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**George Horemis**



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