Read the given statements and select the correct option.

Statement 1: The formula connecting u,v and

f for a spherical mirror is valid only for mirrors whose sizes are very small compared to their radii of curvature.

Statement 2: Laws of reflection are strictly valid for plane surfaces, but not for large spherical surfaces.

- (A) Both statements 1 and 2 are true and statement 2 is the correct explanation of statement 1.
   (B) Both statements 1 and 2 are true but
- statement 2 is not the correct explanation of statement 1. (C) Statement 1 is true but statement 2 is false.
- (D) Statement 1 is false but statement 2 is true.
- 2. A boy throws a table tennis ball of mass 20 g upwards with a velocity of u<sub>0</sub> = 10 m/s at an angle θ<sub>0</sub> with the vertical. The wind imparts a horizontal force of 0.08 N, so that the ball returns to the starting point. Then, the angle θ<sub>0</sub> must be such that, tan θ<sub>0</sub> is

(B) 0.4

(A) 0.2

 A weight is attached to the free end of a sonometer wire. It gives resonance at a length 40 cm when it is resonanced with a tuning fork of frequency 51 Hz. The weight is then immersed

(C) 2.5

(D) 1.2

- wholly in water, the resonant length is reduced to 30 cm. The relative density in which weight suspended is (A) 16/9 (B) 16/7 (C) 16/5 (D) 16/3
- 4. Hydrogen sulphide (H<sub>2</sub>S) contains 94.11% sulphur, water (H<sub>2</sub>O) contains 11.11% hydrogen and sulphur dioxide (SO<sub>2</sub>) contains 50% oxygen. Find the ratio of all given elements. After your calculations which law has been verified?
  (A) Law of multiple proportions
  - (B) Law of reciprocal proportions
     (C) Law of constant composition
     (D) Law of combining volumes
- 5. Which of the following represents the given mode of hybridisation, sp<sup>2</sup>-sp<sup>2</sup>-sp-sp from left to right?

(A) CH<sub>3</sub>=CH−C≡N (B) CH≡C−C≡CH

- (C) CH2=C=C=CH2 (D)
- 6. In the reaction, 4NH<sub>3(g)</sub> + 5O<sub>2(g)</sub> → 4NO<sub>(g)</sub> + 6H<sub>2</sub>O<sub>(f)</sub> when 1 mole of ammonia and 1 mole of O<sub>2</sub> are made to react to completion
  - (A) 1.0 mole of H<sub>2</sub>O is produced
     (B) 2.0 moles of NO will be produced
  - (C) All the oxygen will be consumed
  - (D) All the ammonia will be consumed.

# **National Science Olympiad Sample Papers**

**Sabine Zange** 

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