## THE NATURE OF COVALENT BONDING

## Section Review

## Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used.
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strongth of a covalent bond is related to its bond distociation energy.
- Describe how resonance structures explain bonding.
- Identify some exceptions to the octet rule

## Vocabulary

- · single covalent bond
- structural formulas.
- unshared pairs

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- double covalent bonds
- triple covalent bonds
- coordinate covalent bond
- polyatomic ion
- bond dissociation energy
- resonance structures

#### Part A Completion

tive this completion exercise to check your understanding of the concepts and terms that are Introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the \_\_\_\_\_ configuration of a noble gas, the bonds formed are \_\_ 2\_\_\_ A \_\_ 3 \_\_ pair of valence electrons constitutes a \_\_\_4 \_\_ covalent bond. Pairs of valence electrons that are not shared between atoms are called \_\_\_\_5\_\_. Sometimes two or three pairs of electrons may be shared to give \_\_\_6 \_\_covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bresited attens is known as the \_\_\_9\_\_\_.

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When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a \_\_\_\_10\_\_\_\_

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Eric Strauss, Marylin Lisowski

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