

10.3**THE NATURE OF SOLIDS****SECTION REVIEW****Objectives**

- Describe how the degree of organization of particles distinguishes solids from gases and liquids
- Distinguish between a crystal lattice and a unit cell
- Explain how allotropes of an element differ

Key Terms

- melting point
- crystal
- unit cell
- allotropes
- amorphous solids
- glasses

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Solids tend to be dense and 1. They do not flow or take the shape of their containers, like liquids do, because the particles in solids vibrate around 2 points. When a solid is sufficiently heated and its particles vibrate so strongly that they are no longer held in fixed positions, the solid 3. The 4 is the temperature at which a solid turns to liquid. The melting and 5 of a substance are at the same temperature. In general, ionic solids tend to have relatively 6 melting points while molecular solids tend to have relatively low melting points. Most solids are 7 in nature. The particles are arranged in a pattern known as a crystal 8. The smallest subunit of a crystal lattice is the 9. Some solids have a disordered array of particles and are 10.

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

7. _____

8. _____

9. _____

10. _____

Nature Of Solids Section Review Key

Jin-Ying Zhang



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